



10. Which of the following are endothermic processes?

- A. ice melting
- B. water condensing
- C. freezing
- D. sublimation

11. Heat (enthalpy) is released in EXO, positive temperature change.

12. Temperature following is a statement of the first law of thermodynamics?

- A. is negative change in enthalpy corresponds to an exothermic process
- B. Energy lost by the system must be gained by the surroundings, neither of the two is created
- C. Temperature is a measure of average kinetic energy
- D. Enthalpy is a state function

13. The system enthalpy of a closed process, ΔH is negative. This means that the product is _____

- A. Enthalpically above the reactants, with more energy
- B. Enthalpically above the reactants, with less energy, with disorder
- C. Enthalpically and the surroundings, with more energy
- D. Enthalpically and the surroundings of the surroundings, with more
- E. None of these is a true statement

14. The reaction



is _____ and therefore heat is _____ from the system.

- A. Endothermic, released
- B. Exothermic, released
- C. Endothermic, absorbed
- D. Exothermic, absorbed
- E. Thermodynamic neither released nor absorbed

15. Heat evolved in this reaction when 10.0 g of C_2H_2 is completely consumed in the reaction. How much heat evolved in kJ? $\Delta H = -1300 \text{ kJ}$



- A. 1300 kJ
- B. 650 kJ
- C. 130 kJ
- D. 1300 kJ
- E. 650 kJ



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Jin-Ying Zhang



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