

Year 11 Chemistry Notes - 2021

- Module 1: Properties and Structure of Matter
- Module 2: Introduction to Quantitative Chemistry
- Module 3: Reactive Chemistry
- Module 4: Drivers of Reactions

Module 1: Properties and Structure of Matter **Mixtures**

A mixture is a combination of two or more pure substances (containing one type of molecule) in which each pure substance retains its individual chemical properties - the mixture itself is impure. There are two kinds of mixtures: heterogeneous and homogeneous.

Heterogeneous Mixtures

Two or more substances intermingle, but remain **physically separate**. Often it is possible to separate the original ingredients by **simple physical means**, such as filtering, centrifuge, decanting or sedimentation.

- Examples include: Dirt + Sand, Oil + Water, Salt + Baking Soda
- A Suspension is a specific type of heterogeneous mixture where particles settle at the bottom

Homogeneous Mixtures

Two or more substances have merged into a **uniform phase**. There are no borders between the substances, but they are not chemically bonded. The **physical properties** of each ingredient can be exploited to separate them.

- Examples include: Saltwater, Copper Sulfate solution
- Saltwater can be distilled (boiled) to separate the water
- The two types include solutions and colloids (particles are present, but are very small and do not settle)

Physical properties include magnetism, solubility, density, boiling point, melting point, particle size.

Methods of Separation

Separation Method	Property used	Example
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Annelies Wilder-Smith



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