

1. As a chemical bond forms between two hydrogen atoms in a system, energy is released and the stability of the system

A) decreases **B) increases**
 C) remains the same

2. As two atoms of hydrogen combine to form a molecule of hydrogen, the total energy of the two atoms

A) decreases **B) increases**
 C) remains the same

3. What is conserved during a chemical reaction?

A) mass, only
 B) charge, only
 C) both mass and charge
 D) neither mass nor charge

4. As a chemical bond forms between two hydrogen atoms, the potential energy of the atoms

A) decreases **B) increases**
 C) remains the same

5. As energy is released during the formation of a bond, the stability of the chemical system generally will

A) decrease **B) increase**
 C) remain the same

6. When a valence atom reacts with a chlorine atom to form a compound, the electron configuration of the ion, having the completed *s* and *p* orbitals is closer to that of which noble gas atom?

A) krypton and neon **B) krypton and argon**
 C) neon and helium **D) neon and argon**

7. Which structural diagram represents H_2O ?

A) $\text{H} - \text{O} - \text{H}$ **B) $\text{H} \equiv \text{O}$**
 C) $\begin{array}{c} \text{H} & \text{H} \\ | & | \\ \text{O} & \\ | & | \\ \text{H} & \text{H} \end{array}$ **D) $\begin{array}{c} \text{H} & \text{H} & \text{H} \\ | & | & | \\ \text{O} & & \end{array}$**

8. Which is the correct structural formula for a molecule of chlorine?

A) $\begin{array}{c} \text{Cl} & \text{Cl} \\ | & | \\ \text{Cl} & \text{Cl} \end{array}$ **B) $\text{Cl} - \text{Cl} - \text{Cl} - \text{Cl}$**
 C) $\begin{array}{c} \text{Cl} & \text{Cl} \\ | & | \\ \text{Cl} & \text{Cl} \end{array}$ **D) $\text{Cl} - \text{Cl} - \text{Cl}$**

9. A sodium atom attains a stable electron configuration when it bonds with

A) one chlorine atom **B) two chlorine atoms**
 C) one sodium atom **D) two sodium atoms**

10. Which pair of elements forms a bond with the greatest ionic character?

A) Pb and P **B) Cl and Br**
 C) Na and Cl **D) Cl and N**

11. Which pair of elements forms a bond with the least ionic character?

A) P-Cl **B) Br-Cl**
 C) H-Cl **D) O-Cl**

12. Based on your knowledge of the periodic table, the atoms of which of these elements have the strongest attraction for electrons in a chemical bond?

A) N B) Na C) P **D) Br**

13. As a chlorine atom becomes a negative ion, the atom

A) gains an electron and its radius decreases
 B) gains an electron and its radius increases
 C) loses an electron and its radius increases
 D) loses an electron and its radius decreases

14. Which compound contains ionic bonds?

A) NaCl B) NaOH C) K_2SO_4 **D) CH_4**

15. The bonds in NaCl are best described as

A) covalent, because valence electrons are shared
 B) covalent, because valence electrons are transferred
 C) ionic, because valence electrons are shared
D) ionic, because valence electrons are transferred

16. A substance that has a melting point of 1014°C , conducts electricity when dissolved in water, but does not conduct electricity in the solid phase. The substance is most likely

A) an ionic solid B) a network solid
 C) a molecular solid **D) a metallic solid**

17. Which substance is an electrolyte?

A) CH_3OH **B) $\text{C}_6\text{H}_{12}\text{O}_6$**
 C) H_2O **D) NaOH**

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