

CASCADED MC QUESTIONS: GENERAL CHEMISTRY II

100 Pts (200 Pts Score) (20 Pts per Question)

1. Which among the intermolecular forces is the weakest?
 - A. Hydrogen Bond
 - B. Ion-Dipole Forces
 - C. Dipole-Dipole Forces
 - D. Dispersion Forces
2. In which of the following would dipole-dipole forces exist?
 - A. Butane
 - B. Methane
 - C. CH_2Cl_2
 - D. CH_3OH
3. The vapor pressure of ethanol at 40 °C is 175 mmHg. What would be the vapor pressure at 20 °C?
 - A. 175 mmHg
 - B. Greater than 175 mmHg
 - C. Less than 175 mmHg
 - D. Cannot be determined
4. Which of the following statements about the following process will *not* result in a higher yield?
 - A. All the products of the reaction
 - B. Increased
 - C. It should take some time to reach equilibrium of reaction
 - D. It depends on the nature of the
5. Which of the following would have the lowest boiling point?
 - A. H_2O
 - B. CH_3OH
 - C. H_2
 - D. CH_3CH_3
6. Which of the following reactions would be the most favorable?
 - A. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 - B. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
 - C. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}$
 - D. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{O}_2$
7. Which of the following describes Henry's Law?
 - A. Partial pressure increases with the solubility of gases.
 - B. Partial pressure increases with the solubility of gases.
 - C. Partial pressure increases with the solubility of solids.
 - D. Partial pressure increases with the solubility of solids.
8. How much of the product is formed at 20 °C if 1.00 mol of H_2O is added to 1.00 mol of H_2O ?
 - A. 1.00 mol
 - B. 1.00 g
 - C. 1.00 mol
 - D. 1.00 g
9. Water's vapor pressure at 20 °C is 17.5 mmHg. The temperature is 20 °C. What is the vapor pressure of water at 20 °C?
 - A. 17.5 mmHg
 - B. 17.5 mmHg
 - C. 17.5 mmHg
 - D. 17.5 mmHg
10. Which of the following is the correct statement?
 - A. The vapor pressure of water at 20 °C is 17.5 mmHg. The temperature is 20 °C. What is the vapor pressure of water at 20 °C?
 - B. The vapor pressure of water at 20 °C is 17.5 mmHg. The temperature is 20 °C. What is the vapor pressure of water at 20 °C?
 - C. The vapor pressure of water at 20 °C is 17.5 mmHg. The temperature is 20 °C. What is the vapor pressure of water at 20 °C?
 - D. The vapor pressure of water at 20 °C is 17.5 mmHg. The temperature is 20 °C. What is the vapor pressure of water at 20 °C?

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A Loxley



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