

General Chemistry placement test Study Guide 2024 Solved 100%

Atomic Number - The number of protons in the nucleus of an atom. The atomic number of an element never changes, meaning that the number of protons in the nucleus of every atom in an element is always the same.

Hydrogen = 1, Carbon = 6, Oxygen = 8 (located on top of element of periodic table)

Mass Number - # of neutrons + # of protons = mass number (located on bottom of periodic table)

Isotopes - Let's say an atom is missing a neutron or has an extra neutron. That type of atom is called an isotope. An atom is still the same element if it is missing an electron.

Example: (carbon-14)

Distribution &/W subatomic particles in an Atom - A typical atom consists of three subatomic particles: Protons, Neutrons, and Electrons.

- Nucleons include protons and neutrons. All the positive charge of an atom is contained in the nucleus, and originates from the protons. Neutrons are neutrally-charged. Electrons, which are negatively-charged, are located outside of the nucleus.

Roles of subatomic particles: Protons, Neutrons, & Electrons - **Particle Mass:**

Protons & Neutrons have about same mass. Electrons have much less mass than protons and neutrons.

Particle Charge:

Protons have a positive charge (+)

- Neutrons have no charge (0)

- Net charge of the nucleus is positive and equal to the number of protons.

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Xiaolong Qi



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Table of Contents General Chemistry Assessment Test Study Guide

1. Understanding the eBook General Chemistry Assessment Test Study Guide
 - The Rise of Digital Reading General Chemistry Assessment Test Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Assessment Test Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Assessment Test Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Assessment Test Study Guide
 - Personalized Recommendations
 - General Chemistry Assessment Test Study Guide User Reviews and Ratings
 - General Chemistry Assessment Test Study Guide and Bestseller Lists
5. Accessing General Chemistry Assessment Test Study Guide Free and Paid eBooks

- General Chemistry Assessment Test Study Guide Public Domain eBooks
- General Chemistry Assessment Test Study Guide eBook Subscription Services
- General Chemistry Assessment Test Study Guide Budget-Friendly Options
- 6. Navigating General Chemistry Assessment Test Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Assessment Test Study Guide Compatibility with Devices
 - General Chemistry Assessment Test Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Assessment Test Study Guide
 - Highlighting and Note-Taking General Chemistry Assessment Test Study Guide
 - Interactive Elements General Chemistry Assessment Test Study Guide
- 8. Staying Engaged with General Chemistry Assessment Test Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Assessment Test Study Guide
- 9. Balancing eBooks and Physical Books General Chemistry Assessment Test Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Assessment Test Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Assessment Test Study Guide
 - Setting Reading Goals General Chemistry Assessment Test Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Assessment Test Study Guide
 - Fact-Checking eBook Content of General Chemistry Assessment Test Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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