

100) Carbonic acid reacts as a buffer at  $\text{pH}$  7.4. How many moles of  $\text{H}_2\text{CO}_3$  are needed to buffer 100 mL of 0.1 M  $\text{NaHCO}_3$  solution?

$$\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$$

$$7.4 = 6.35 + \log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]}$$

$$1.05 = \log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]}$$

$$10^{1.05} = \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]}$$

$$11.22 = \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]}$$

$$[\text{H}_2\text{CO}_3] = \frac{[\text{HCO}_3^-]}{11.22}$$

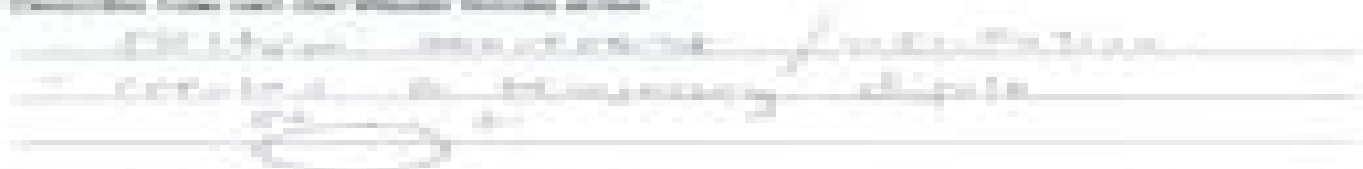
$$[\text{H}_2\text{CO}_3] = \frac{0.1 \text{ M}}{11.22}$$

$$[\text{H}_2\text{CO}_3] = 0.0089 \text{ M}$$

moles =  $0.0089 \text{ M} \times 0.1 \text{ L} = 0.00089 \text{ moles}$  100

101) What are the two main types of buffer systems in the blood?

101a) Bicarbonate buffer system and phosphate buffer system.



101b) Bicarbonate buffer system is the most important buffer system in the blood. 100

101c) Bicarbonate buffer system is the most important buffer system in the blood.



102) Bicarbonate buffer system is the most important buffer system in the blood.

102a) Bicarbonate buffer system is the most important buffer system in the blood.

102b) Bicarbonate buffer system is the most important buffer system in the blood.

moles =  $0.0089 \text{ M} \times 0.1 \text{ L} = 0.00089 \text{ moles}$  100



# F321 Past Papers 2006

**Sabine Zange**



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