| 11 to | Which compound he composition by mass | as the greatest percent | 1.8 | Which unit is used to thermal energy? |
|--------------|------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| The state of | (1) BaS (2) CaS | (3) MgS (4) SrS | | (1) gram (2) mole |
| UP! | Two molecules of HB | is released is absorbed | 19 Judy PP | Under which condition pressure does a real gaideal gas? (1) low temperature and (2) low temperature and (4) high temperature and (4) |
| Second . | electron configuratio | The state of the s | 20 Way | According to the kinetic ideal gas, all gas particles |
| 本州 | (1) earbon (2) magnesium | (4) oxygen | #18 | (2) are in random, consta (2) are separated by very to their sizes |
| | bond? | scribes a <u>multip</u> le covalent | | (3) have strong intermole (4) have collisions that de of the system |
| May | (1) Two electrons as (2) Four electrons as (3) Two electrons as (4) Four electrons as | re shared. e transferred. | 21 | Which mathematical exp heat of reaction for a che |
| Most | The electronegativi atoms in a molecul determine | by difference between the | way to | (1) (the heat of fusion) – (2) (the heat of vaporization) (the potential energy potential energy of the potential energy potential energy of the poten |
| 451 | (d) the polarity of atoms | on energy of the atoms the bond between the two an not be broken down by | 22 122 4 17 | At 101.3 kPa and 298 K which compound absorb of heat as the entire sam |
| Many | chemical means? (1) CO and He (2) CO and NH ₃ | ③ Xe and He Table (4) Xe and NHa S | | (2) NaCl + 3.38 For a reaction at equilib increase the rates of the reactions? |
| 17 | (2) chemically comb | a mixture differ in density ese properties can be used to stances sine the substances reezing point of the mixture etrical conductivity of the | | (1) a decrease in the reactants (2) a decrease in the surf an increase in the term of the forward reaction |

- express an amount of Table
 - (4) pascal
- (3) joule
- is of temperature and is behave most like an
- low pressure
 - high pressure
 - d low pressure
 - d high pressure
 - molecular theory for an
- mt, straight-line motion small distances relative
 - scular forces
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 - the heat of vaporization)
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 - rium, which change cun e forward and reverse
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 - see area of the products
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