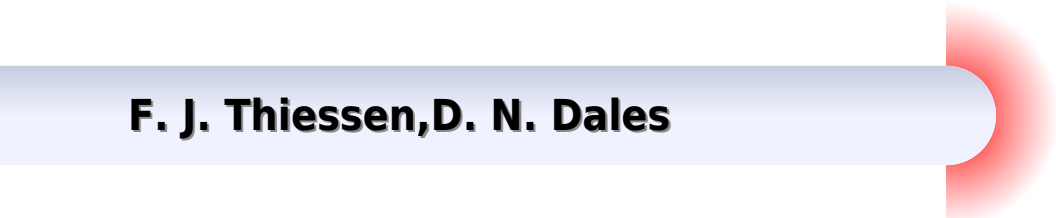




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F. J. Thiessen, D. N. Dales



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web equilibrium moles $hscn$ $2.00 \times 10^{-5} \text{ mol}$ $3.00 \times 10^{-6} \text{ mol}$ $1.70 \times 10^{-5} \text{ mol}$ $hscn$ knowing the number of moles of Fe^{3+} and $hscn$ present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species $mol\ Fe^{3+}$ $1.70 \times 10^{-5} \text{ mol}$ Fe^{3+} $8.50 \times 10^{-4} \text{ M}$

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zeroes are 3 and 4 is $x^2 + 12x + 12$ b $x^2 + 12x - 12$ c $x^2 - 12x + 12$ d $x^2 - 12x - 12$

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is the zero of the polynomial we know that if α is a zero of the polynomial $p(x)$ then $p(\alpha) = 0$ substituting $x = 2$ in $x^2 + 3x + p$
 $2^2 + 3(2) + p = 0$
 $4 + 6 + p = 0$
 $10 + p = 0$
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