

HISTORY GRADE 11 (CAPS)
TABLE OF PRESCRIBED AND EXAMINABLE CONTENT - PAPERS 1 AND 2 (2014)

PRESCRIBED CONTENT	EXAMINABLE CONTENT	MID YEAR EXAMINATION	
		PAPER 1	EXAM STRUCTURE
TOPIC 1 : Communism in Russia 1900-1940	Lenin's interpretation of Marxism or Stalin's interpretation of Marxism-Leninism	Source-based Question	ONE three (3) hour paper Learners must answer THREE (3) questions One Essay Question (50) + One Source-based question (50) + One other question (50) Total marks = 150
TOPIC 2 : Capitalism in USA 1900- 1940	The February and October Revolution of 1917	<u>Essay Question</u>	
	Capitalism boom of the 1920s, USA society in 1920s, wall street crash of 1929.	Source-based Question	
TOPIC 3: Ideas of Race in the 19th and 20th centuries	Analysis and assessment of the New Deal	<u>Essay Question</u>	
	Case study of Nazi Germany and the Holocaust	Source-Based Question	
	Theory and Practise of Eugenics (examples from any case study can be used)	<u>Essay Question</u>	
PRESCRIBED CONTENT	EXAMINABLE CONTENT	END OF YEAR EXAMINATION	
		PAPER 2	Exam Structure
TOPIC 4: Nationalisms – South Africa	The influence of WWII on the rise of African nationalism	Source-based Question	TWO x three (3) hour papers Learners must answer THREE (3) questions One Essay Question (50) + One Source-based question (50) + One other question (50) P1 Total marks = 150 P2 Total marks = 150
TOPIC 4: Nationalisms – Middle East and Africa	The rise of Afrikaner nationalism	<u>Essay Question</u>	
	The 1948 war – The Palestinian and Israeli perspectives	Source-Based Question	
TOPIC 5: Apartheid South Africa 1940s to 1960s	The influence of WWII on the rise of nationalism in 'Gold Coast' (Ghana)	<u>Essay Question</u>	
	Segregation in the 1920s and 1930s.	Source-Based Question	
	Resistance to Apartheid before 1960.	<u>Essay Question</u>	

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$ if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have $\frac{21.3 \text{ g } \text{O}_2}{32 \text{ g/mol}} = 0.6656 \text{ mol}$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g/mol} = 302.3 \text{ g}$ c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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