CHI SELECT AND A WARRENCE - Force A

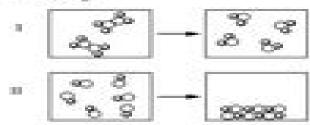
Fill in your name, CID and section on this test booklet. Fill in and bubble in your name, CID (for the letter "C" bubble zero), section and "DI" on the scantron form. For question 966, mark A. For each question, read all possible choices then mark the best one on your scantron. Use a #2 possil and make clean evaners to insure proper scoring. When finished, turn in your scantron and show this booklet and your CID to the proctor. All questions are worth 4 points each, Let your knowledge show?

1. Which of the following is balanced correctly?

A aboveled equation is balanced when the same manufact of attents are in the reactions and products. The reactions is B contain 16 angger attents, but the products and products. The reactions is B contain 16 percentage attents in C postales a contain and 12 pB; The reactions in C postales a contain a discount in B contain 6 contain attents, 16 Rydrogen attents, and 8 anggers attents. The products contain a contain 2 contain 2 contain 2 contains.

 Which of the following is an evidation-reduction reaction?

 Which of the following transfers heat to the autropulings?



- At A confer
- III. III mode
- Brooth Based III.
- Discontinue I or III

En Piguer E, the hand between the two white spheres in breaken. In Piguer II, we made with combine to force a well-I proving that. Found broading cognitive contragtembrica, while hand forming releases energy transferences.

 How many grams of N.H; would be required to react completely with 99.0 g of O₂ according to the belanced chemical position?

$$2 \text{ NoH}_1(g) = 3 \text{ Octob} \rightarrow 4 \text{ NO}_1(g) + 2 \text{ HoOcto}$$

A. 377.1 g.

D 232 g

B. 42.1 gt

E. 268 a

- 40. 24.2 回
- The electrical expectation is used to relate the number of molice of moles of another. Of greats of Or is given, it mouls to be commented to moles. The coefficients are used as mole material, and the formula moles of NAL is used to get buck to get

on
$$\phi_H \circ_A \left(\frac{3 \mod O_1}{3 \log O_2}\right) \left(\frac{2 \mod N_1 H_1}{3 \mod O_1}\right) \left(\frac{3 \log N_1 N_2}{3 \mod O_1}\right)$$

= 37.5 g N.H.

General Chemistry Lab Answers

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