

Chemistry 105 Equation Sheet

Constants

$$\begin{aligned}a_0 &= 0.5291772059 \text{ \AA} \\c &= 2.99792458 \times 10^8 \text{ m/s} \\e &= 1.6021765 \times 10^{-19} \text{ C} \\\epsilon_0 &= 8.8541877 \times 10^{-12} \text{ F/m} \\\hbar &= 6.626 \times 10^{-34} \text{ J}\cdot\text{s} \\k_B &= 1.38065 \times 10^{-23} \text{ J/K} \\k_B &= 8.987551 \times 10^8 \frac{\text{J}\cdot\text{m}^2}{\text{mol}\cdot\text{K}\cdot\text{C}^2} \\m_e &= 9.109383 \times 10^{-31} \text{ kg} \\m_n &= 1.6749273 \times 10^{-27} \text{ kg} \\m_p &= 1.6726217 \times 10^{-27} \text{ kg} \\N_A &= 6.022142 \times 10^{23} \text{ mol}^{-1} \\R &= 8.31447 \frac{\text{J}}{\text{mol}\cdot\text{K}} \\&= 0.0820575 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} \\&= 1.987207 \frac{\text{cal}}{\text{mol}\cdot\text{K}} \\R_H &= 1.0974 \times 10^7 \text{ m}^{-1}\end{aligned}$$

Conversions (* = exact)

Length

$$\begin{aligned}1 \text{ in} &= 2.54 \text{ cm}^* \\1 \text{ ft} &= 12 \text{ inches}^* \\1 \text{ yd} &= 3 \text{ ft}^* \\1 \text{ m} &= 39.37 \text{ inches}^* \\1 \text{ mile} &= 5280 \text{ feet}^* \\1 \text{ mile} &= 1.609 \text{ km}\end{aligned}$$

Mass

$$\begin{aligned}1 \text{ lb} &= 453.59237 \text{ g} \\&= 16 \text{ oz. avdp.}^* \\1 \text{ US lb} &= 12 \text{ troy ounces}^* \\1 \text{ ton} &= 2000 \text{ lb}^* \\1 \text{ amu} &= 1.660538 \times 10^{-24} \text{ g}\end{aligned}$$

Volume

$$\begin{aligned}1 \text{ US gal} &= 3.7854 \text{ L} = 4 \text{ US qt}^* \\1 \text{ US qt} &= 32 \text{ US fl. oz.}^* \\1 \text{ US fluid oz.} &= 29.5735 \text{ mL} \\1 \text{ cm}^3 &= 0.001 \text{ L}^* \\1 \text{ in}^3 &= 16.4 \text{ cm}^3\end{aligned}$$

Pressure

$$\begin{aligned}1 \text{ atm} &= 760 \text{ torr}^* \\&= 101325 \text{ Pa}^* \\&= 406.783 \text{ in H}_2\text{O} \\&= 14.6959 \text{ lb/in}^2 \text{ (psi)} \\1 \text{ torr} &= 1 \text{ mm Hg (at } 0^\circ\text{C)}^* \\1 \text{ Pa} &= 1 \text{ kg}\cdot\text{m}^{-1}\cdot\text{s}^{-2} \\1 \text{ bar} &= 100,000 \text{ Pa}^*\end{aligned}$$

Force

$$1 \text{ N} = 1 \text{ kg}\cdot\text{m}\cdot\text{s}^{-2}$$

Power

$$1 \text{ Watt} = 1 \text{ J/s}$$

Energy

$$\begin{aligned}1 \text{ J} &= 1 \text{ kg}\cdot\text{m}^2/\text{s}^2 = 0.2390\text{-cal} \\1 \text{ J} &= 1 \text{ V}\cdot\text{C} = 0.009369 \text{ L}\cdot\text{atm} \\1 \text{ eV} &= 1.602 \times 10^{-19} \text{ J} \\1 \text{ L}\cdot\text{atm} &= 101.3 \text{ J}\end{aligned}$$

Formulas

$$\begin{aligned}\rho &= m/V \\T_F &= \frac{9}{5}T_C + 32^\circ\text{F} \\T_K &= \frac{1}{1}T_C + 273.15 \text{ K} \\c &= \lambda \nu \\E &= h\nu \\\lambda &= \frac{h}{mv} \\E_{\text{kinetic}} &= \frac{1}{2}mv^2 \\E &= mc^2 \\\Phi &= h\nu - KE_{\text{electron}} \\E_n &= -Z^2 \frac{R_H h c}{n^2}\end{aligned}$$

$$\Delta x \Delta(mv) \geq \frac{h}{4\pi}$$

$$\frac{1}{\lambda} = R_H \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$V(r) = \frac{q_1 q_2}{4\pi\epsilon_0 r}$$

$$F(r) = k_e \frac{Q_1 Q_2}{r^2}$$

$$\mu = Qr$$

$$\Delta E = q + w$$

$$\Delta E_{\text{univ}} = 0$$

$$\Delta S_{\text{univ}} \geq 0$$

$$S = k_B \ln W$$

$$q = mC\Delta T$$

$$w = -P_{\text{ext}} \Delta V$$

$$\Delta H = \Delta E + P\Delta V$$

$$\Delta S = \frac{q_{\text{rev}}}{T}$$

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta G^\circ + RT \ln Q$$

$$\Delta H'_{\text{rxn}} = \sum nH'_{f,\text{prod}} - \sum nH'_{f,\text{react}}$$

$$\Delta H'_{\text{rxn}} \approx \sum nD_{\text{broken}} - \sum nD_{\text{formed}}$$

$$PV = nRT$$

$$P_{\text{tot}} = P_1 + P_2 + P_3 + \dots$$

$$\frac{\text{Rate}_A}{\text{Rate}_B} = \frac{\text{Time}_B}{\text{Time}_A} = \sqrt{\frac{M_B}{M_A}}$$

$$KE_{\text{ave}} = \frac{3}{2}RT$$

$$u_{\text{rms}} = \sqrt{v^2} = \sqrt{\frac{3RT}{M}}$$

$$pX = -\log X$$

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-]$$

* = exact ν = frequency v = velocity

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General Chemistry Formula And Equations Sheet

Clifford Lane



General Chemistry Formula And Equations Sheet:

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