

09/08/2020

# ## Chapter 1: General Chem Review - Part 1 ##

- 1.1 Reactions: a) molecules collide  
b) bonds broken/formed

Organic compounds  $\rightarrow$  compounds w carbon atoms

- 1.2 Structural Theory: substances are defined by a specific arrangement  
compounds differ based on the way the atoms are bonded

Constitutional isomer: same molecular formula but different structures

Atoms commonly bonded to C: N, O, H & halides

tetravalent  
(4 bonds)

trivalent  
(3 bonds)

divalent  
(2 bonds)

monovalent  
(one bond)

- 1.3 Covalent bond: pair of  $e^-$  shared by two atoms

Energy released when forming bond  $\rightarrow$  equilibrium internuclear distance

internuclear repulsion if contracted further  $\rightarrow$  (bond length)

$\uparrow \uparrow$  space  $= \downarrow$  energy of electrons

energy @ equilibrium  
is bond strength

Potential Energy & stability:  $\downarrow$  PE = most stable

What keeps bond @ optimal length?

+ nuclei & - electron  $\rightarrow$  attractive forces

+ nuclei & + nuclei  $\rightarrow$  repulsive forces  $\leftarrow$  - electron & - electron

Lewis Structure Examples:



- 1.4 Formal charge: valence electrons it has vs. valence electrons need to be neutral  
eg. CH3O



C has 4 valence electrons  $\rightarrow$  no formal charge  
Needs 4 more to be neutral

O has 7 electrons  $\rightarrow$  1- formal charge  
Needs 1 less to be neutral

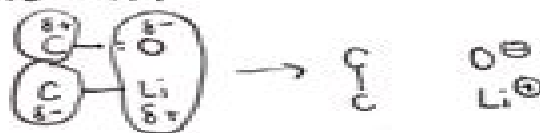
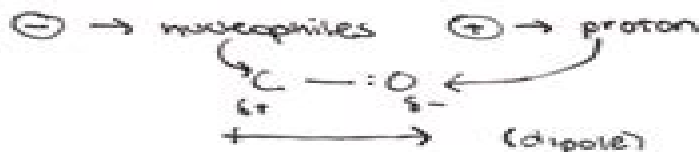
- 1.5 Electronegativity: increase L to R & B to T  
F is most EN

$\uparrow$  carbon usually forms polar covalent bonds

Covalent bond: EN difference  $< 0.5$

Polar covalent bond: EN difference  $0.5 - 1.7$

Ionic bond: EN difference  $> 1.7$



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