

100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of H_2CO_3 are needed to buffer 1.0 mole of HCO_3^- at pH 7.4?

$$\text{H}_2\text{CO}_3 \rightleftharpoons \text{H}^+ + \text{HCO}_3^- \quad \text{pK}_a = 6.35$$

$$\text{pH} = 7.4$$

$$\text{pH} = \text{pK}_a + \log \frac{[\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3]}$$

101) The only intracellular buffers in body fluids are not the phosphate.

102) Describe how each of the following buffers work:

1. Phosphate buffer system
acts as a primary buffer
in the cytoplasm
in the mitochondria
in the blood

103) Describe how each of the following buffers work:

1. Protein buffer system
acts as a secondary buffer
in the cytoplasm
in the blood

104) The Henderson-Hasselbalch equation is a logarithmic relationship between the concentration of a weak acid and its conjugate base.

105) The Henderson-Hasselbalch equation is a logarithmic relationship between the concentration of a weak acid and its conjugate base.

106) The Henderson-Hasselbalch equation is a logarithmic relationship between the concentration of a weak acid and its conjugate base.



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PT Brinkman



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