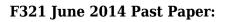
| Abbreviations. | 1 | = alternative and acceptable answers for the same marking point |
|-----------------|-------|--|
| annotations and | 10000 | = separates marking points |
| conventions | NOT | = answers which are not worthy of credit |
| used in the | () | = words which are not essential to gain credit |
| | 200 | (underlining) key words which <u>must</u> be used to gain credit |
| Mark Scheme | ecf. | = error carried forward |
| | AW | = alternative wording |
| | ora | = or reverse argument |

| Question | Expected answers | Marks |
|----------|---|-------|
| 1 (a) | 8 1, 2 or 3 p orbitals are OK | [2] |
| (b) | d orbital 2 ✓ p sub-shell 6 ✓ 3rd shell 18 ✓ | [3] |
| | energy $2s$ $\uparrow \downarrow$ $2p$ $\uparrow \downarrow \uparrow$ \uparrow \uparrow \downarrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \uparrow \uparrow \downarrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \downarrow \uparrow \downarrow \uparrow \downarrow \uparrow \downarrow \downarrow \downarrow \downarrow \uparrow \downarrow | |
| | 2s and 2p labels ✓ Ignore any superscripted numbers. 8 electrons in correct levels with arrows correctly shown ✓ | [2] |
| (d) (| Energy change when each atom in 1 mole ✓ of gaseous atoms ✓ loses an electron ✓ (to form 1 mole of gaseous 1+ ions). | [3] |
| (| O²⁺(g) → O³⁺(g) + e⁻ ✓ ✓ 1 mark for correct species; 1 mark for state symbols. No charge required on electron. | [2] |
| | Ignore (g) on e | 121 |

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Emilie Sanchez



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