

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "n-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

H ⁺		Transition elements								N ³⁻	O ²⁻	F ¹⁻	
Li ⁺	Be ²⁺							Al ³⁺		P ³⁻	S ²⁻	Cl ¹⁻	
Na ⁺	Mg ²⁺												
K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺								Br ³⁻
Rb ⁺	Sr ²⁺			Ag ¹⁺				Sn ²⁺ Sn ⁴⁺					I ¹⁻
	Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺			Pb ²⁺ Pb ⁴⁺					
Cations										Anions			

1. Based on the information in Model 1:

a. Identify three elements that form only one cation.

$$\text{Cu}^+ \quad \text{K}^+ \quad \text{Na}^+$$

b. Identify three elements that form only one anion.

+ charged element

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