

## Ionic Compounds & Properties

Ionic compounds typically exist as non-volatile solids due to the strong ionic bonds holding the lattice together. This means that the ions are stuck in their positions in the lattice and are not free to move. However, when dissolved in water, the water molecules surround the ions and separate them from each other; this means that the ions are now free to move. An electrical current is simply the flow of charge, so ionic compounds can conduct electricity when dissolved or molten.

## Metallic Bonding & Alloys

If we consider elemental metals or alloys, there is only one type of atom - metals. Metals have low electronegativity, but also low differences in electronegativity - This is how a metallic bond is defined. The bonding in metals is between the positive nuclei and the sea of delocalized electrons. This means that the strength of the bond may be determined by the charge density of the cation and the number of electrons that are delocalized over the lattice. Alloys are typically mixtures of two or more metals (and possibly carbon) and have properties proportional to the ratio of the metals. The properties depend on the properties and lattice structure of the constituent atoms.

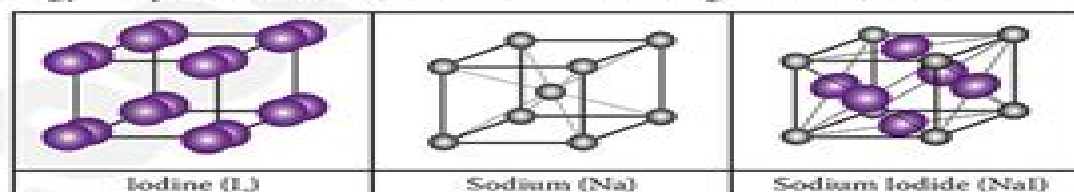
## Melting Point

Melting points are a simple way of estimating the relative strength of bonds between whatever is making up the lattice. Typically giant structures have high melting points because of the strong forces holding the lattice together. Even so, we can use melting points to describe the trends in the strength of the various types of bonds.

The melting point is the temperature that gives the particles enough kinetic energy so that they can overcome the attractive forces acting in the lattice of the solid crystal.

Melting points are a measure of the attractive forces that hold the lattice together.

Melting does not decompose a compound into its elements - it is only a change of state! Many students confuse the overcoming of intermolecular forces with decomposition. In most cases, a covalent molecule will change state by overcoming intermolecular forces, well before it decomposes by breaking the stronger covalent bonds that hold the molecule together. Consider Iodine,  $I_2$ , the energy of vaporization is  $41 \text{ kJ mol}^{-1}$  vs. the bond strength of  $151 \text{ kJ mol}^{-1}$ .



## Melting points of Group I – The alkali metals

Before we discuss the melting point, we must recall the type of bonding – in this case metallic.

Metallic bonding is the attractive force between the lattice of positive ions, and the “sea” of delocalized electrons. There are three factors affecting the strength of the attraction:

- 1) – The charge on the ion (eg. +1, +2, +3 for Na, Mg, and Al respectively)
- 2) – The size of the ion. – as the ions increases in size, the charge density decreases.
- 3) – The number of delocalized electrons (the opposite charge as the ion)

Know it!

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