

# Ionic Chemical Formulas/PreAICE Chemistry

NAME: \_\_\_\_\_ PERIOD \_\_\_\_\_

**Directions:** Using the Chemical Formulas program sheet, fill in the table below. Use subscripts & superscripts. Remember to use roman numerals to show the valence of the transition metals in the compound name.

METAL CATION (+)	CATION FORMULA	NONMETAL ANION (-)	ANION FORMULA	COMPOUND FORMULA	COMPOUND NAME
sodium	Na <sup>+1</sup>	hydroxide	OH <sup>-</sup>	NaOH	sodium hydroxide
magnesium	Mg <sup>+2</sup>	phosphate	PO <sub>4</sub> <sup>-3</sup>	Mg <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	magnesium phosphate
calcium	Ca <sup>+2</sup>	chloride	Cl <sup>-</sup>	CaCl <sub>2</sub>	calcium chloride
gold(I)	Au <sup>+1</sup>	sulfite	SO <sub>3</sub> <sup>-2</sup>	Au <sub>2</sub> SO <sub>3</sub>	gold(I) sulfite
copper(II)	Cu <sup>+2</sup>	hydroxide	OH <sup>-</sup>	Cu(OH) <sub>2</sub>	copper(II) hydroxide
Iron(III)	Fe <sup>+3</sup>	telluride	Te <sup>-2</sup>	Fe <sub>2</sub> Te <sub>3</sub>	iron(III) telluride
magnesium	Mg <sup>+2</sup>	bromide	Br <sup>-1</sup>	MgBr <sub>2</sub>	magnesium bromide
manganese(II)	Mn <sup>+2</sup>	sulfide	S <sup>-2</sup>	MnS	manganese(II) sulfide
boron	B <sup>+3</sup>	oxide	O <sup>-2</sup>	B <sub>2</sub> O <sub>3</sub>	boron oxide
potassium	K <sup>+</sup>	nitrate	NO <sub>3</sub> <sup>-</sup>	KNO <sub>3</sub>	potassium nitrate
barium	Ba <sup>+2</sup>	bromide	Br <sup>-</sup>	BaBr <sub>2</sub>	barium bromide
strontium	Sr <sup>+2</sup>	sulfate	SO <sub>4</sub> <sup>-2</sup>	SrSO <sub>4</sub>	strontium sulfate
iron(III)	Fe <sup>+3</sup>	oxide	O <sup>-2</sup>	Fe <sub>2</sub> O <sub>3</sub>	iron(III) oxide
silver(I)	Ag <sup>+1</sup>	chloride	Cl <sup>-</sup>	AgCl	silver(I) chloride
aluminum	Al <sup>+3</sup>	carbonate	CO <sub>3</sub> <sup>-2</sup>	Al <sub>2</sub> (CO <sub>3</sub> ) <sub>3</sub>	aluminum carbonate
gold(II)	Au <sup>+2</sup>	nitride	N <sup>-3</sup>	Au <sub>3</sub> N <sub>2</sub>	gold(II) nitride
cesium	Cs <sup>+1</sup>	phosphate	PO <sub>4</sub> <sup>-3</sup>	Cs <sub>3</sub> PO <sub>4</sub>	cesium phosphate
ammonium	NH <sub>4</sub> <sup>+</sup>	sulfide	S <sup>-2</sup>	(NH <sub>4</sub> ) <sub>2</sub> S	ammonium sulfide
zinc(II)	Zn <sup>+2</sup>	nitrite	NO <sub>2</sub> <sup>-</sup>	Zn(NO <sub>2</sub> ) <sub>2</sub>	zinc(II) nitrite
chromium(I)	Cr <sup>+1</sup>	nitride	N <sup>-3</sup>	Cr <sub>3</sub> N	chromium(I) nitride
mercury(II)	Hg <sup>+2</sup>	oxide	O <sup>-2</sup>	HgO	mercury(II) oxide
sodium	Na <sup>+</sup>	sulfite	SO <sub>3</sub> <sup>-2</sup>	Na <sub>2</sub> SO <sub>3</sub>	sodium sulfite
magnesium	Mg <sup>+2</sup>	oxide	O <sup>-2</sup>	MgO	magnesium oxide

# Ions Answer Key

**Yijin Wang**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name Yijin Wang.

## **Ions Answer Key:**

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*The Structure of the Mercury-solution Electrical Double Layer in the Presence of Adsorbed Oxyanions* Richard Payne,1966 The unusual adsorption characteristics of oxyanions are described with special reference to the adsorption of nitrate ions from  $\text{KNO}_3$  solutions The adsorption energy is shown to be linearly

dependent on the charge in contrast to the behavior previously found for mixed solutions of  $\text{NH}_4\text{NO}_3$  and also  $\text{NH}_4\text{ClO}_4$  with  $\text{NH}_4\text{F}$  at constant ionic strength where the dependence is approximately quadratic

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This book delves into Ions Answer Key. Ions Answer Key is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ions Answer Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Ions Answer Key
- Chapter 2: Essential Elements of Ions Answer Key
- Chapter 3: Ions Answer Key in Everyday Life
- Chapter 4: Ions Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Ions Answer Key. This chapter will explore what Ions Answer Key is, why Ions Answer Key is vital, and how to effectively learn about Ions Answer Key.
3. In chapter 2, this book will delve into the foundational concepts of Ions Answer Key. The second chapter will elucidate the essential principles that must be understood to grasp Ions Answer Key in its entirety.
4. In chapter 3, the author will examine the practical applications of Ions Answer Key in daily life. This chapter will showcase real-world examples of how Ions Answer Key can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Ions Answer Key in specific contexts. The fourth chapter will explore how Ions Answer Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Ions Answer Key. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Ions Answer Key.

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## **Table of Contents Ions Answer Key**

1. Understanding the eBook Ions Answer Key

- The Rise of Digital Reading Ions Answer Key
- Advantages of eBooks Over Traditional Books
- 2. Identifying Ions Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ions Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Ions Answer Key
  - Personalized Recommendations
  - Ions Answer Key User Reviews and Ratings
  - Ions Answer Key and Bestseller Lists
- 5. Accessing Ions Answer Key Free and Paid eBooks
  - Ions Answer Key Public Domain eBooks
  - Ions Answer Key eBook Subscription Services
  - Ions Answer Key Budget-Friendly Options
- 6. Navigating Ions Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Ions Answer Key Compatibility with Devices
  - Ions Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ions Answer Key
  - Highlighting and Note-Taking Ions Answer Key
  - Interactive Elements Ions Answer Key
- 8. Staying Engaged with Ions Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ions Answer Key

9. Balancing eBooks and Physical Books Ions Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ions Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ions Answer Key
  - Setting Reading Goals Ions Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ions Answer Key
  - Fact-Checking eBook Content of Ions Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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