

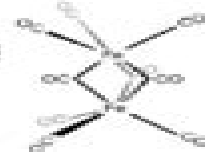
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## CHEM 2218 Inorganic Chemistry I (Final Exam sample paper)

### Part I: Multiple Choices (2 points each, 50 points in total)

- For a cube, the number of  $C_6$ ,  $C_4$ , and  $C_2$  rotational axes is  
(a) 1, 3, and 4, respectively (b) 0, 2, and 4, respectively (c) 0, 3, and 4, respectively
- The CO diatomic molecule contains infinite number of  
(a)  $\sigma_v$  (b)  $i$  (c)  $\sigma_h$  (d)  $C_2$
- Identify the wrong label(s) in the following list used for representations of point groups.  
 $C_{1000}$ ,  $C_{100}$ ,  $D_{1000}$ ,  $D_{2v}$ ,  $C_4$ ,  $I$ ,  $O_h$ ,  $T_d$
- The  $Fe_2(CO)_9$  molecule shown does not possess what element of symmetry?  
(a)  $i$  (b)  $\sigma$  (c)  $S_6$  (d)  $C_2$
- A molecule having a point group of  $D_{6h}$  contains all the symmetry elements of a molecule having a point group of  
(a)  $D_{3d}$  (b)  $D_{2d}$  (c)  $D_{3h}$  (d) none of the above
- How many mirror planes of symmetry are there in an icosahedron?  
(a) 6 (b) 9 (c) 12 (d) 15 (e) 18
- Give the symmetry label for the molecular orbital associated with a trigonal planar structure (*you are given a character table for  $D_{3h}$* ).  
(a)  $A_1^+$  (b)  $A_1^-$   
(c)  $A_2^+$  (d)  $A_2^-$
- Identify the stronger acid?  
 $H_2Se$  vs.  $H_2S$        $[Al(H_2O)_6]^{3+}$  vs.  $[Fe(H_2O)_6]^{3+}$
- Which one in the following is the least stable?  
(a)  $[Al(H_2O)_6]^{3+}$  (b)  $[AlF_6]^{3-}$  (c)  $Mo(H_2O)_6$  (d)  $Mo(CO)_6$
- Predict which way the following reactions will go (left or right)  
(a)  $[AlF_6]^{3-} + Ga(CN)_6]^{3-} \rightarrow [GaF_6]^{3-} + Al(CN)_6]^{3-}$   
(a)  $W(CO)_5(PMe_3) + NMe_3 \rightarrow W(CO)_5(NMe_3) + PMe_3$
- Select the strongest base toward proton.  
(a)  $NH_3$  (b)  $NHMe_2$  (c)  $PH_3$  (d)  $PMe_3$
- Which of the following reactions belong to the acid-base reactions?  
(a)  $H + OH \rightarrow H_2O$ .  
(b)  $PF_3 + F_2 \rightarrow PF_5$ .  
(c)  $SiF_4 + 2NaF \rightarrow Na_2[SiF_6]$ .  
(d)  $Mo + 6CO \rightarrow Mo(CO)_6$ .



# Inorganic Chemistry Question Papers

**L Cohen**



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