



CHEMISTRY

CHEAT SHEET



NAMING SIMPLE COMPOUNDS

NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Their Formulas
(Show the name corresponding to the formula)



Directions for naming:

1. Name the cation first, then the anion.
2. For ionic compounds, the cation name is the same as the element name.
3. For covalent compounds, the first element is named as is, and the second element is named with a prefix and an -ide suffix.
4. The order of the elements in the formula is the same as the order in the name.
5. The prefix for the number of atoms of each element is shown in parentheses.
6. The name of the compound is written with the elements in the same order as the formula.

Binary Compounds

First Element	Anion
H ₂	hydrogen
Li	lithium
Na	sodium
K	potassium
Rb	rubidium
Cs	cesium
F	fluorine
Cl	chlorine
Br	bromine
I	iodine
S	sulfur
O	oxygen
N	nitrogen
P	phosphorus
As	arsenic
Sb	antimony
Bi	bismuth

Common Ions (aq)

Formula	Name
H ⁺	hydrogen
Li ⁺	lithium
Na ⁺	sodium
K ⁺	potassium
Rb ⁺	rubidium
Cs ⁺	cesium
F ⁻	fluoride
Cl ⁻	chloride
Br ⁻	bromide
I ⁻	iodide
S ²⁻	sulfide
O ²⁻	oxide
N ³⁻	nitride
P ³⁻	phosphide
As ³⁻	arsenide
Sb ³⁻	antimonide
Bi ³⁻	bismuthide

Naming Compounds with a Central and Terminal



Carbon Dioxide

Directions for naming:

1. Name the central element first, then the terminal element.
2. The name of the central element is the same as the element name.
3. The name of the terminal element is the same as the element name, but with a prefix and an -ide suffix.
4. The order of the elements in the formula is the same as the order in the name.

Common Ions

Formula	Name	Formula	Name
H ⁺	hydrogen	Li ⁺	lithium
Li ⁺	lithium	Na ⁺	sodium
Na ⁺	sodium	K ⁺	potassium
K ⁺	potassium	Rb ⁺	rubidium
Rb ⁺	rubidium	Cs ⁺	cesium
Cs ⁺	cesium	F ⁻	fluoride
F ⁻	fluoride	Cl ⁻	chloride
Cl ⁻	chloride	Br ⁻	bromide
Br ⁻	bromide	I ⁻	iodide
I ⁻	iodide	S ²⁻	sulfide
S ²⁻	sulfide	O ²⁻	oxide
O ²⁻	oxide	N ³⁻	nitride
N ³⁻	nitride	P ³⁻	phosphide
P ³⁻	phosphide	As ³⁻	arsenide
As ³⁻	arsenide	Sb ³⁻	antimonide
Sb ³⁻	antimonide	Bi ³⁻	bismuthide
Bi ³⁻	bismuthide		

Common Ions for Specific Elements

Formula	Name	Formula	Name
H ⁺	hydrogen	Li ⁺	lithium
Li ⁺	lithium	Na ⁺	sodium
Na ⁺	sodium	K ⁺	potassium
K ⁺	potassium	Rb ⁺	rubidium
Rb ⁺	rubidium	Cs ⁺	cesium
Cs ⁺	cesium	F ⁻	fluoride
F ⁻	fluoride	Cl ⁻	chloride
Cl ⁻	chloride	Br ⁻	bromide
Br ⁻	bromide	I ⁻	iodide
I ⁻	iodide	S ²⁻	sulfide
S ²⁻	sulfide	O ²⁻	oxide
O ²⁻	oxide	N ³⁻	nitride
N ³⁻	nitride	P ³⁻	phosphide
P ³⁻	phosphide	As ³⁻	arsenide
As ³⁻	arsenide	Sb ³⁻	antimonide
Sb ³⁻	antimonide	Bi ³⁻	bismuthide
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Naming Compounds Study Guide

David R. Klein



Naming Compounds Study Guide:

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02

Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011

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Basic Concepts of Chemistry, Study Guide Leo J. Malone, 1989-02-21 This Third Edition revised to provide smoother transitions between topics employs a concise yet informal approach to basic chemistry organized to help students employ basic math skills and problem solving strategies Writing style is straightforward and presentation incorporates many concrete analogies to clarify new concepts Includes many illustrative worked examples

Students Study Guide Fredeen, Pearson, 2003-07

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