

- (T) A student estimated quantitatively x g of $(\text{NH}_4)_2\text{SO}_4$. Excess NaOH C, containing 1.00dm^3 of 2.05mol dm^{-3} was heated with the $(\text{NH}_4)_2\text{SO}_4$ to expel all the NH_3 gas. 25.0cm^3 portion of excess solution C after treatment with $(\text{NH}_4)_2\text{SO}_4$ was titrated with 0.195mol dm^{-3} H_2SO_4 using few drops of phenolphthalein as indicator. The table below was presented as follows:

Burette readings (cm^3)	1st	2nd	3rd
Final burette reading	25.05	23.55	31.05
Initial burette reading	1.50	2.70	10.80
Volume of H_2SO_4 used	20.55	20.85	20.25

FASTANSWER TEAM

The equation of the reaction between NaOH solution and H_2SO_4 solution is:



[Na = 23.0; O = 16.0; H = 1.00; S = 32.0; N = 14.0]

- a. (i) Mention the laboratory apparatus used to measure the 25.0cm^3 portion of solution C and state why it should be used?

- (ii) State whether phenolphthalein indicator is suitable or unsuitable. Give reason for the answer.

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- (iii) State why few drops of phenolphthalein indicator was used.

- (iv) Explain briefly why the student needed to repeat another titration exercise.

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- (v) State three experimental errors that might be committed by the student.

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