

## Section A

Answer all questions in the spaces provided

0 1

Anhydrous magnesium chloride,  $\text{MgCl}_2$ , can absorb water to form the hydrated salt  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}$



← creating crystals directly is not really possible.

0 1 - 1

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

0 1 - 2

Some enthalpies of solution are shown in Table 1.

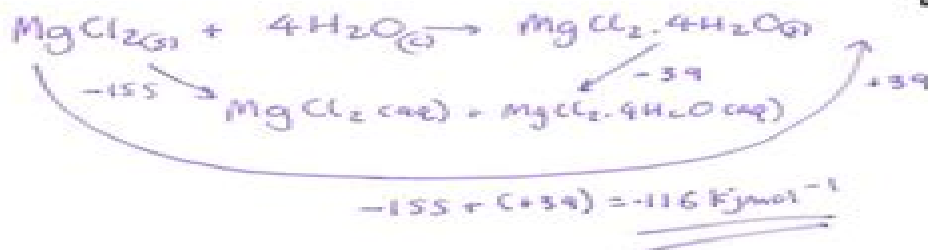
Table 1

Salt	Enthalpy of solution / $\text{kJ mol}^{-1}$
$\text{MgCl}_2(\text{s})$	-155
$\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$	-39

making the salts aqueous!

Calculate the enthalpy change for the absorption of water by  $\text{MgCl}_2(\text{s})$  to form  $\text{MgCl}_2 \cdot 4\text{H}_2\text{O}(\text{s})$ .

[2 marks]



Enthalpy change \_\_\_\_\_  $\text{kJ mol}^{-1}$



# Max Papers Chemistry Alevel 2014

**Michael Seilmaier**



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**Cambridge International AS and A Level Chemistry** Peter Cann, Peter Hughes, 2015-03-06 Endorsed by Cambridge Assessment International Education for full syllabus coverage Foster a deeper understanding of theoretical concepts through clear guidance and opportunities for self assessment throughout covers the entire Cambridge International AS A Level Chemistry syllabus 9701 Navigate the different routes through the course with ease with clearly divided sections for AS and A Level Focus learning with learning outcomes clearly defined at the beginning of each section Test knowledge and understanding with past paper and exam style questions Address the Key Concepts in the syllabus which are clearly highlighted throughout the course The Revision and Practice CD included with every Student's Book provides interactive tests summaries of each topic and advice on examination techniques

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