

Chemical Reactions of Copper: Key

Copper Recycling Lab

(Day 1)

Day 1 Data

1. Write the balanced chemical equation for the reaction between copper and nitric acid (HNO_3). For all reactants and products, include the state of matter.



Cancel spectators: NO_3^-



Copper Recycling Lab

(Day 2 & 3)

1. What evidence was there that a chemical reaction took place in conversion 1?

Formation of $\text{Cu(NO}_3)_2$ solution (blue), release of NO_2 (g) brown, heat change and condensation

2. Identify the type of chemical reaction for conversion 1.

Single Displacement

3. What was the limiting reactant in conversion 1 and how can you tell?

Cu (s) - All of the copper wire is used up

4. Classify the product in the flask after conversion 1 as a pure substance, homogeneous mixture, or heterogeneous mixture.

Homogeneous Mixture

5. Why is the copper (II) nitrate not visible once the reactions are done? Where did it go?

It is an aqueous solution - is dissolved

6. If the experiment stopped after conversion 1, what method could be used to recover this compound?

Evaporate off the water

Mass In Chemical Reactions Lab Answers

Julia Schneider



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