

All answers in cells and lines are valued at 1 pt.
each unless marked otherwise.

Name _____ Period _____

Atoms and Bonds Chapter Test

Section 1 Fill in the Table

Compound/Molecule	Metal + Nonmetal or Nonmetal + Nonmetal?	Ionic or Covalent Bond?	e ⁻ shared or transferred?
KI			
Br ₂			
CO ₂			
PbF ₂			

Section 2 Periodic Symbol Analysis

28
Ni
58.69

Element name _____
 Name English or Latin in origin? _____
 Atomic number _____ Number of protons _____
 Atomic mass _____ Number of electrons _____
 Number of neutrons (Mass - # protons) _____
 Metal, nonmetal, or noble gas? _____

Section 3 Properties of Bonds Write "I" if the property describes an ionic bond, "C" if the property describes a covalent bond, and "B" if the property describes both kinds of bonds.

_____ High melting points _____ Weaker bond _____ Electrons shared
 _____ Conducts electricity in water _____ Helps most atoms achieve the octet of a noble gas
 _____ Polar or nonpolar molecules _____ Attraction of opposite charges
 _____ Poor conductors of electricity _____ Crystal shape _____ Low melting points
 _____ Attraction between a cation and an anion

Section 4 Formula Writing Write the formula for each of the compounds below. Be sure to balance the charges with subscripts. (Hint: Parenthesized Roman numerals are the charges of transition metals)

- Silver sulfide _____
- Nickel (II) cyanide _____
- Aluminum phosphate _____
- Aluminum bromide _____
- Copper (I) sulfate _____
- Silver bromide _____

Cyanide = CN⁻ Sulfide = S²⁻ Phosphate = PO₄³⁻ Silver = Ag⁺ Sulfate = SO₄²⁻ Aluminum = Al³⁺

Ionic Bonds Chapter Test Answer Key

Yi-Tong Ma



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Table of Contents Ionic Bonds Chapter Test Answer Key

1. Understanding the eBook Ionic Bonds Chapter Test Answer Key
 - The Rise of Digital Reading Ionic Bonds Chapter Test Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bonds Chapter Test Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bonds Chapter Test Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bonds Chapter Test Answer Key
 - Personalized Recommendations
 - Ionic Bonds Chapter Test Answer Key User Reviews and Ratings
 - Ionic Bonds Chapter Test Answer Key and Bestseller Lists

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 - Ionic Bonds Chapter Test Answer Key Public Domain eBooks
 - Ionic Bonds Chapter Test Answer Key eBook Subscription Services
 - Ionic Bonds Chapter Test Answer Key Budget-Friendly Options
6. Navigating Ionic Bonds Chapter Test Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bonds Chapter Test Answer Key Compatibility with Devices
 - Ionic Bonds Chapter Test Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bonds Chapter Test Answer Key
 - Highlighting and Note-Taking Ionic Bonds Chapter Test Answer Key
 - Interactive Elements Ionic Bonds Chapter Test Answer Key
8. Staying Engaged with Ionic Bonds Chapter Test Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bonds Chapter Test Answer Key
9. Balancing eBooks and Physical Books Ionic Bonds Chapter Test Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bonds Chapter Test Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ionic Bonds Chapter Test Answer Key
 - Setting Reading Goals Ionic Bonds Chapter Test Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Bonds Chapter Test Answer Key
 - Fact-Checking eBook Content of Ionic Bonds Chapter Test Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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