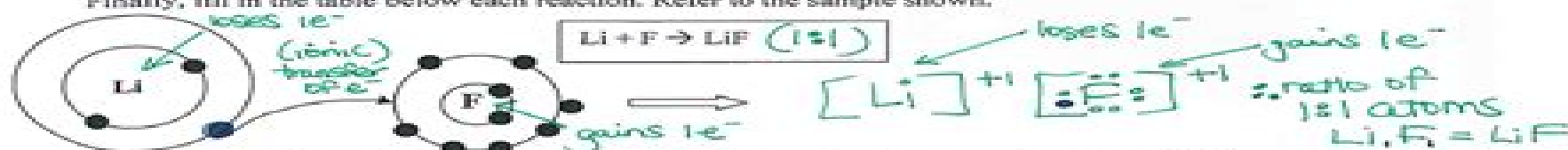


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li \cdot + \cdot \ddot{Cl} : \rightarrow [Li]^+ [\ddot{Cl}:]^{-1}$	Li	1 ve	donates $1e^-$	Li^+
	Cl	1 ve	accepts $1e^-$	Cl^{-1}
2) $Ca + O \Rightarrow CaO$ $Ca \cdot + \cdot \ddot{O} : \rightarrow [Ca]^{+2} [\ddot{O}:]^{-2}$	Ca	2 ve	donates $2e^-$	Ca^{+2}
	O	6 ve	accepts $2e^-$	O^{-2}
3) $Be + F \Rightarrow BeF_2$ $Be \cdot + \cdot \ddot{F} : \rightarrow [Be]^{+2} [\ddot{F}:]^{-1}$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
	F	7 ve	each F accepts $1e^-$ each	F^{-1}
4) $Mg + S \Rightarrow MgS$ $Mg \cdot + \cdot \ddot{S} : \rightarrow [Mg]^{+2} [\ddot{S}:]^{-2}$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
	S	6 ve	accepts $2e^-$	S^{-2}
5) $K + F \Rightarrow KF$ $K \cdot + \cdot \ddot{F} : \rightarrow [K]^+ [\ddot{F}:]^{-1}$	K	1 ve	donates $1e^-$	K^{+1}
	F	7 ve	accepts $1e^-$	F^{-1}

Ionic Bond Worksheet Answer Key

JA Banks



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Table of Contents Ionic Bond Worksheet Answer Key

1. Understanding the eBook Ionic Bond Worksheet Answer Key
 - The Rise of Digital Reading Ionic Bond Worksheet Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Bond Worksheet Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ionic Bond Worksheet Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Bond Worksheet Answer Key
 - Personalized Recommendations

- Ionic Bond Worksheet Answer Key User Reviews and Ratings
- Ionic Bond Worksheet Answer Key and Bestseller Lists
- 5. Accessing Ionic Bond Worksheet Answer Key Free and Paid eBooks
 - Ionic Bond Worksheet Answer Key Public Domain eBooks
 - Ionic Bond Worksheet Answer Key eBook Subscription Services
 - Ionic Bond Worksheet Answer Key Budget-Friendly Options
- 6. Navigating Ionic Bond Worksheet Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ionic Bond Worksheet Answer Key Compatibility with Devices
 - Ionic Bond Worksheet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ionic Bond Worksheet Answer Key
 - Highlighting and Note-Taking Ionic Bond Worksheet Answer Key
 - Interactive Elements Ionic Bond Worksheet Answer Key
- 8. Staying Engaged with Ionic Bond Worksheet Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ionic Bond Worksheet Answer Key
- 9. Balancing eBooks and Physical Books Ionic Bond Worksheet Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ionic Bond Worksheet Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ionic Bond Worksheet Answer Key
 - Setting Reading Goals Ionic Bond Worksheet Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ionic Bond Worksheet Answer Key
 - Fact-Checking eBook Content of Ionic Bond Worksheet Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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