

Conclusion

Through gravimetric analysis, we were able to determine the percentage of chloride in the dried sample. One of the most important sources of error in this experiment is the nature of the precipitate. The result of adding the precipitate in a hasty manner instead of doing it slowly will be ended up form a colloidal precipitate which was the reason why solution in the beaker does not readily distinguish the separation of the solvent to precipitate. Besides that, Using high concentration of substance that was measured and weight and precipitant can also lead to form of not good precipitate or it take longer than usual to get a good result. To achieve good precipitate, one must constantly stir the solution while heating because it is important to digest the colloidal sample. Other than that, the sample may absorb moisture in the air when it was cooling down may lead to an increasing weight compared to the original. It is important to note about systematic error when carry out the experiment. Nevertheless, it proved that gravimetric analysis can accurately determine the result of the experiment as much as stoichiometric method.

Recommendation

- Make sure the experiment must be carry out in one go.
- Keep stirring the solution while heating to avoid coagulation.
- The crucible must be completely dried before weighting it to avoid measuring error.

Reference

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Appendices

1. Equation: $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \longrightarrow \text{NaCl}(\text{aq})$

How To Write A Conclusion For A Lab Report

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Table of Contents How To Write A Conclusion For A Lab Report

1. Understanding the eBook How To Write A Conclusion For A Lab Report
 - The Rise of Digital Reading How To Write A Conclusion For A Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying How To Write A Conclusion For A Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an How To Write A Conclusion For A Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from How To Write A Conclusion For A Lab Report

- Personalized Recommendations
- How To Write A Conclusion For A Lab Report User Reviews and Ratings
- How To Write A Conclusion For A Lab Report and Bestseller Lists
- 5. Accessing How To Write A Conclusion For A Lab Report Free and Paid eBooks
 - How To Write A Conclusion For A Lab Report Public Domain eBooks
 - How To Write A Conclusion For A Lab Report eBook Subscription Services
 - How To Write A Conclusion For A Lab Report Budget-Friendly Options
- 6. Navigating How To Write A Conclusion For A Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - How To Write A Conclusion For A Lab Report Compatibility with Devices
 - How To Write A Conclusion For A Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of How To Write A Conclusion For A Lab Report
 - Highlighting and Note-Taking How To Write A Conclusion For A Lab Report
 - Interactive Elements How To Write A Conclusion For A Lab Report
- 8. Staying Engaged with How To Write A Conclusion For A Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers How To Write A Conclusion For A Lab Report
- 9. Balancing eBooks and Physical Books How To Write A Conclusion For A Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection How To Write A Conclusion For A Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine How To Write A Conclusion For A Lab Report
 - Setting Reading Goals How To Write A Conclusion For A Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of How To Write A Conclusion For A Lab Report

- Fact-Checking eBook Content of How To Write A Conclusion For A Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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