Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point.

Express the heat in kilojoules to three significant figures.

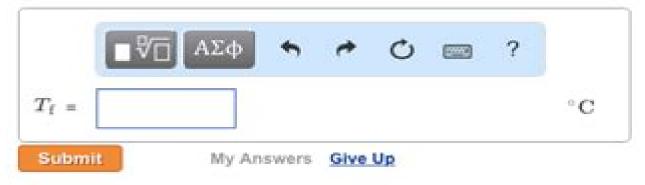
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At its boiling point, the enthalpy of vaporization ($\Delta H_{\rm vap}$) of water is 40.7 kJ/mol. Once the given mass, 2.58×10^3 g, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by $\Delta H_{\rm vap}$ yields the amount of heat required to evaporate that amount of water.

Part B

Suppose that 0.48 g of water at 25° C condenses on the surface of a 55-g block of aluminum that is initially at 25° C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum, $C_{s,A1}$, is 0.903 $J/(g \cdot {}^{\circ}C)$.)

Express the temperature in degrees Celsius to two significant figures.



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