



## OWNER'S MANUAL

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**⚠ CAUTION**

Read this manual before operating or servicing snowplow.

This document supersedes all editions with an earlier date.

# Manual For Suburbanite Western Plow

**Zhe Xu**



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web a rate law shows how the rate of a chemical reaction depends on reactant concentration for a reaction such as  $a \rightarrow \text{products}$  the rate law generally has the form  $\text{rate} = k [a]^n$  where  $k$  is a proportionality constant called the rate constant and  $n$  is the order of the reaction with respect to  $a$  the value of  $n$  is not related to the reaction stoichiometry and

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web the rate determining slower step gives a rate law showing second order dependence on the  $\text{no } 2$  concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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to be directly proportional to the concentration of  $n$  2o 5 text rate  $k$  n

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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