NAME:		

HYDROCARBON NAMING WORKSHEET

- Naming (Given the structure be able to write the name)
 - Number longest carbon chain
 - Write the main group
 - Name and number the substituents (Alphabetically)
 - A prefix of (di, tri, tetra) is used when there are more than two of same substituent

Naming Breakdown					
Numbers	two or more of same substituent	Substituents	Parent	endings	
1	di	methyl	meth	-ane	
3	tri	ethyl	eth	-ene	
3	tetra	propyl	prop	-yme	
4		butyl	but	benzene	
5		fluoro	pent		
6 7		chiloro	hex		
7		bromo	hept		
8		iodo	oct		
9		phenyl	non	52	
10		nitro	dec		
		sulfonic			
		cyclo	5		

1. Given the name draw the structure. Identify each 1°, 2°, 3°, or 4° carbon of an alkane.

a) 3-methylheptane

$$\sim$$

b) 1-chloro-3-ethylcyclopentane



c) 2, 3-dichlorohexane

Hydrocarbon Study Guide Answer Sheet

Gary J. Blomquist, Anne-Geneviève Bagnères

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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011 A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013 A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject guickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text Student Guide For Living Chemistry Barbara Ucko, 2012-12-02 Student Guide for Living Chemistry is a 23 chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit Each objective corresponds to a section in the textbook Living Chemistry Three kinds of questions are provided for each objective to check the student's understanding namely short answer Study Questions multiple choice and fill in The answers for all guestions are provided at the end of the chapter The opening chapters cover the SI units composition of matter chemical bonding compounds chemical change gases respiration and water The subsequent chapters deal with solutions acids bases salts nuclear and organic chemistry oxygen derivatives and hydrocarbons polymers and other organic derivatives This textbook also explores the chemistry of carbohydrates lipids proteins enzymes and energy and carbohydrate metabolism The remaining chapters discuss the chemistry of vitamins hormones body fluid drugs and poisons Undergraduate chemistry students will find this book invaluable Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry Katherine J. Denniston, Denniston Katherine, 2006 Chemistry, Student **Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex

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