

- Carbohydrates contain carbon, hydrogen and oxygen, usually with the general formula $C_nH_{2n}O_n$

Glucose

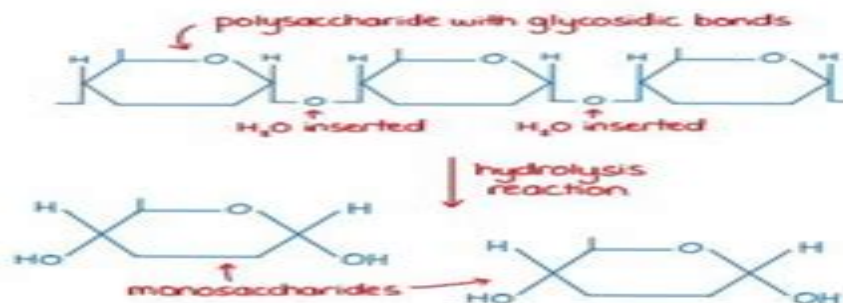
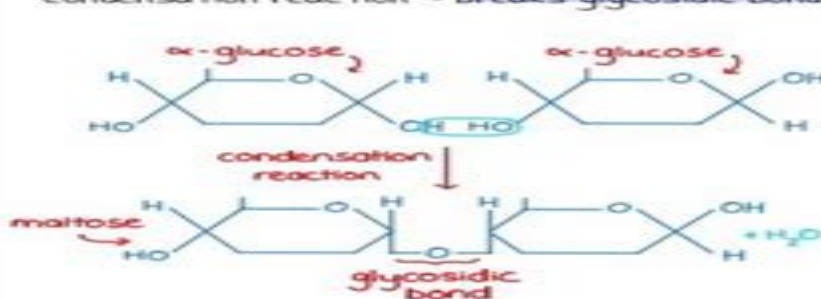
- A monosaccharide with the formula $C_6H_{12}O_6$
- A hexose monosaccharide (six carbon atoms) in a ring structure
- Soluble in water → easily transported
- Main energy source for animals and plants
→ chemical bonds store lots of energy
- Two isomers: α -glucose and β -glucose → H and OH groups on carbon 1 inverted in β -glucose

Monosaccharides are small soluble carbohydrate monomers. They also include fructose and galactose.



Glycosidic bonds and condensation/hydrolysis reactions

- Condensation reaction: two molecules join to form a new chemical bond and a water molecule is eliminated
- Condensation reactions form glycosidic bonds between monosaccharides to create disaccharides and polysaccharides
- Hydrolysis reaction: a water molecule is used and the chemical bond is broken (reverse of a condensation reaction → breaks glycosidic bonds)



Disaccharides

- Two monosaccharides joined together with a glycosidic bond in a condensation reaction
- Soluble in water
- Maltose = glucose + glucose
- Sucrose = glucose + fructose
- Lactose = glucose + galactose

Monosaccharides and disaccharides are sugars.

Polysaccharides

- Large polymers of monosaccharides joined with glycosidic bonds
- Starch and glycogen are large energy storage molecules which cannot leave cells

Starch

- Glucose storage in plants → hydrolysed when glucose is needed
- Insoluble in water → does not affect the water potential of cells so water is not drawn in by osmosis
- Amylose → unbranched α -glucose polysaccharide (1,4-glycosidic bonds)

Iodine test for starch

- Add iodine in potassium iodide solution to sample.
- If starch is present: goes from brown/orange to blue-black

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Glycolysis Revision A Revision Note

CO Houle



Glycolysis Revision A Revision Note:

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Table of Contents Glycolysis Revision A Revision Note

1. Understanding the eBook Glycolysis Revision A Revision Note
 - The Rise of Digital Reading Glycolysis Revision A Revision Note
 - Advantages of eBooks Over Traditional Books
2. Identifying Glycolysis Revision A Revision Note
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Glycolysis Revision A Revision Note
 - User-Friendly Interface
4. Exploring eBook Recommendations from Glycolysis Revision A Revision Note
 - Personalized Recommendations
 - Glycolysis Revision A Revision Note User Reviews and Ratings

- Glycolysis Revision A Revision Note and Bestseller Lists
- 5. Accessing Glycolysis Revision A Revision Note Free and Paid eBooks
 - Glycolysis Revision A Revision Note Public Domain eBooks
 - Glycolysis Revision A Revision Note eBook Subscription Services
 - Glycolysis Revision A Revision Note Budget-Friendly Options
- 6. Navigating Glycolysis Revision A Revision Note eBook Formats
 - ePub, PDF, MOBI, and More
 - Glycolysis Revision A Revision Note Compatibility with Devices
 - Glycolysis Revision A Revision Note Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Glycolysis Revision A Revision Note
 - Highlighting and Note-Taking Glycolysis Revision A Revision Note
 - Interactive Elements Glycolysis Revision A Revision Note
- 8. Staying Engaged with Glycolysis Revision A Revision Note
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Glycolysis Revision A Revision Note
- 9. Balancing eBooks and Physical Books Glycolysis Revision A Revision Note
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Glycolysis Revision A Revision Note
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Glycolysis Revision A Revision Note
 - Setting Reading Goals Glycolysis Revision A Revision Note
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Glycolysis Revision A Revision Note
 - Fact-Checking eBook Content of Glycolysis Revision A Revision Note
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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