

Question 1

- 1(b)(i) Reaction rate decreased as time increased.
- 1(b)(ii) The smaller zinc particles gave a larger surface area which led to a faster reaction rate since there were more successful collisions per unit time.
- 1(b)(iii) Temperature, pressure (for gases) and concentration (of reagents in solution).
- 1(c) One safety precaution could have been the wearing of safety glasses/goggles to protect the eyes from possible acid spray.
- 1(d) Temperature could have been controlled during the experiment by the use of a water bath.
- 1(e)(i) $\text{Zn}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{ZnSO}_{4(aq)} + \text{H}_{2(g)}$
- 1(e)(ii) From the balanced equation for the reaction, 1 mole of $\text{Zn}_{(s)}$ produces 1 mole of $\text{H}_{2(g)}$.
 Moles of zinc used = mass of Zn used/molar mass of Zn = $(0.55/65.4)$ mol, i.e. 0.0084 mol. Zn.
 Therefore 0.0084 moles of $\text{H}_{2(g)}$ were produced.
 Volume of $\text{H}_{2(g)}$ produced = moles of $\text{H}_{2(g)}$ produced $\times$ molar volume of gas at RTP.
 Volume of $\text{H}_{2(g)}$ produced = $(0.0084 \times 24\,000) \text{ cm}^3 = 201.8 \text{ cm}^3$.

Question 2

- 2(a) The number of occupied electron shells in an tells us what period it is in. All these elements have three occupied electron shells and are therefore in period three.

- 2(b) Element: Sodium (Na).



- 2(c) Allotrope – one of multiple (2 or more) different forms/structures of the same element, which exist in the same physical state.

- 2(d)  
 × – (sulfur) electron.
 S – nucleus of sulfur atom.

- 2(e)(i) When sulfur reacts with chlorine there is covalent bonding.

- 2(e)(ii) Both sulfur and chlorine are non-metals. It is energetically preferable for these atoms to share electrons with each other in order to achieve stable outer electronic configurations of 8 electrons, rather than to lose or gain electrons.

- 2(f) SCl_2 .

- 2(g) 
 S – sulfur nucleus and inner electrons, × – sulfur electron,
 Cl – chlorine nucleus and inner electrons, • – chlorine electron.

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Jianjun Gao



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 - The Rise of Digital Reading May 2014 Chemistry Cxc Past Paper Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying May 2014 Chemistry Cxc Past Paper Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an May 2014 Chemistry Cxc Past Paper Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from May 2014 Chemistry Cxc Past Paper Answers
 - Personalized Recommendations
 - May 2014 Chemistry Cxc Past Paper Answers User Reviews and Ratings
 - May 2014 Chemistry Cxc Past Paper Answers and Bestseller Lists
5. Accessing May 2014 Chemistry Cxc Past Paper Answers Free and Paid eBooks
 - May 2014 Chemistry Cxc Past Paper Answers Public Domain eBooks
 - May 2014 Chemistry Cxc Past Paper Answers eBook Subscription Services
 - May 2014 Chemistry Cxc Past Paper Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - May 2014 Chemistry Cxc Past Paper Answers Compatibility with Devices
 - May 2014 Chemistry Cxc Past Paper Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of May 2014 Chemistry Cxc Past Paper Answers
 - Highlighting and Note-Taking May 2014 Chemistry Cxc Past Paper Answers
 - Interactive Elements May 2014 Chemistry Cxc Past Paper Answers
8. Staying Engaged with May 2014 Chemistry Cxc Past Paper Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers May 2014 Chemistry Cxc Past Paper Answers
- 9. Balancing eBooks and Physical Books May 2014 Chemistry Cxc Past Paper Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection May 2014 Chemistry Cxc Past Paper Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine May 2014 Chemistry Cxc Past Paper Answers
 - Setting Reading Goals May 2014 Chemistry Cxc Past Paper Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of May 2014 Chemistry Cxc Past Paper Answers
 - Fact-Checking eBook Content of May 2014 Chemistry Cxc Past Paper Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Interactive and Gamified eBooks

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