

CHEM 113 Final – Monday May 16th 10:15-12:15 in LSC 133; bring a pencil!

Atomic Structure:

1. What do these have in common?



- A) The same number of protons
- B) The same number of electrons
- C) The same number of neutrons
- D) The same size

2. The element X occurs natural to the extent of 20.0% ^{12}X and 80.0% ^{13}X . The atomic mass of X is nearest

- A) 12.2
- B) 12.5
- C) 12.8
- D) 13.0

3. What is the valence electron configuration for the element in Period 5, Group 3A?

- A) $5s^25p^1$
- B) $3s^23p^3$
- C) $3s^23p^5$
- D) $5s^25p^3$

4. What is the maximum number of electrons that can occupy an orbital labeled d_{xy} ?

- A) 1
- B) 2
- C) 4
- D) 10

5. Which electron transition in a hydrogen atom is associated with the largest emission of energy?

- A) $n = 2$ to $n = 1$
- B) $n = 1$ to $n = 2$
- C) $n = 2$ to $n = 4$
- D) $n = 4$ to $n = 2$

Molecular Structure and Bonding:

1. Which molecule has exactly two unshared (lone) pairs of electrons on the central atom?

- A) BF_3
- B) OF_2
- C) NF_3
- D) XeF_2

2. The molecule of the type ML_4 consists of four single bonds and no lone pairs. What structure is it expected to assume?

- A) square planar
- B) trigonal pyramidal
- C) trigonal planar
- D) tetrahedral

Mastering Chemistry Chem 113 Exercise Answers

Mark Steyvers



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