

1. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temp was found to have  $[\text{CO}] = 0.115\text{M}$ ,  $[\text{H}_2] = 0.110\text{M}$ , and  $[\text{CH}_3\text{OH}] = 0.190\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}][\text{H}_2]^2} = \frac{(0.19)}{(0.115)(0.11)^2} = 136.5$$

2. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temperature was found to have  $[\text{NH}_3] = 0.274\text{M}$  and  $[\text{H}_2\text{S}] = 0.355\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = [\text{NH}_3][\text{H}_2\text{S}] = (0.274)(0.355) = 0.0973$$

3. Consider the following reaction:



The reaction between CO and  $\text{H}_2$  is carried out at a specific temperature with initial concentrations of  $\text{CO} = 0.32\text{M}$  and  $\text{H}_2 = 0.45\text{M}$ . At equilibrium, the concentration of  $\text{CH}_3\text{OH}$  is  $0.14\text{M}$ . Find the equilibrium constant at this temperature.

	CO	$2\text{H}_2$	$\text{CH}_3\text{OH}$
I	0.32	0.45	
C	-x	-2x	+x
E	0.18	0.17	0.14

$$K_c = \frac{(0.14)}{(0.18)(0.17)^2} = 26.9$$

4. Consider the following reaction:



The reaction of  $\text{CH}_4$  is carried out at some temperature with an initial concentration of  $[\text{CH}_4] = 0.090\text{M}$ . At equilibrium, the concentration of  $\text{H}_2$  is  $0.012\text{M}$ . Find the equilibrium constant at this temperature.

	$2\text{CH}_4$	$\text{C}_2\text{H}_2$	$3\text{H}_2$
I	0.09		
C	-2x	+x	+3x
E	0.082	0.004	0.012M

$$K_c = \frac{(0.004)(0.012)^3}{(0.082)^2} = 1.0 \times 10^{-6}$$

5. Consider the following reaction:



A reaction mixture contains  $[\text{NO}_2] = 0.0248\text{M}$  and  $[\text{N}_2\text{O}_4] = 0.0337\text{M}$ . Calculate  $Q_c$  at the given concentrations.

$$Q_c = \frac{(0.0248)^2}{(0.0337)} = 0.0183 \quad Q_c > K_c, \text{ rxn goes left}$$

6. The following reaction was carried out in a 2.75L reaction vessel at  $1100\text{K}$ :



# Mastering Chemistry Answers Chapter 19

**Morris Hein, Susan Arena, Cary Willard**



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**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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computer interaction and online learning this book is ideally designed for academics researchers school administrators professionals and practitioners interested in the design of optimized educational technologies

**Mastering Organic Chemistry** Gerald W. Gibson, 1979 Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation

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