Name: AP Chemistry
Period: Dute: Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

- 5. Please do #18 in chapter 12 of your text.
  - a. This increases the [H<sub>2</sub>], which will increase the rate, but has no effect on k.
  - b. Due to Arrhenius equation, changing temperature changes the value of k.
  - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of k.
- Consider the reaction: P<sub>4</sub> + 6 H<sub>2</sub> → 4 PH<sub>3</sub>. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$\{P_4\}$ , mot/L	[H <sub>2</sub> ], mol/L	Initial Rate, mol/(L : s)
0.0110	0.0075	3.20 × 10 <sup>-4</sup>
0.0110	0.0150	6.40 × 10 <sup>-4</sup>
0.0220	0.0150	6.39 x 10 <sup>-4</sup>

- What is the order with respect to: P<sub>4</sub> \_\_\_\_\_\_\_\_.
  - H<sub>2</sub> \_\_l\_\_\_
- b. Write the rate law for this reaction. rate = k[H<sub>2</sub>]
- c. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp't 1 and solving for k, we get k = 0.0427 s<sup>-1</sup>
- Consider the reaction: SO<sub>2</sub> + O<sub>3</sub> → SO<sub>3</sub> + O<sub>2</sub>. A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

[SO <sub>2</sub> ], mol/L	{O <sub>3</sub> }, mol/L	Initial Rate, mol/(L s)
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- a. What is the order with respect to: SO<sub>2</sub> \_\_2\_.
  O<sub>1</sub> O<sub>2</sub>
- b. Write the rate law for this reaction. rate = k[SO-]<sup>3</sup>[O-]<sup>4</sup>
- c. Determine the value and units of the rate constant, k. plug and chug using the rate law & data from exp't 1 and solving for k, we get k = 2.36 mol L<sup>-1</sup> s<sup>-1</sup>
- 8. Consider the following mechanism.  $A_2 + B_2 \rightarrow R + C$  (slow)
  - $A_2 + R \rightarrow C$  (fast)
  - a. Write the overall balanced chemical equation.  $2A_2 + B_2 \rightarrow 2C$
  - b. Identify any intermediates within the mechanism. R
  - c. What is the order with respect to each reactant? A2 I"; B2 I"

# **Mastering Chemistry Answer Key Kinetics**

**Siddharth Venkatesh** 

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connection with HDN mechanisms although in a much more modest scale and depth Some years ago when encouraged by Prof B R James this book was planned several excellent reviews and monographs treating different aspects of HDS were already available including some on the subject of organometallic models However it seemed appropriate to try to summarize the most striking features of this chemistry in an updated and systematic way and inasmuch as possible in connection with the common knowledge and beliefs of the mechanisms of heterogeneous HDS catalysis Hopefully this attempt to build some conceptual bridges between these two traditionally separated areas of chemistry has met with some success Chemistry Siddharth Venkatesh, 2025-02-20 Discover the essential aspects of chemistry in various industries with Applied Chemistry Practical Applications This comprehensive textbook provides an in depth understanding of fundamental chemical principles and their real world applications Covering a wide range of topics from chemical reactions and materials science to environmental chemistry and sustainable practices it caters to students researchers and professionals Written by experts our book blends theoretical concepts with practical examples offering a solid foundation in key concepts followed by discussions on their applications in industry technology and everyday life We emphasize sustainability green chemistry principles and environmentally friendly practices Clear explanations of complex topics are supported by diagrams illustrations and tables Our book integrates modern research findings and technological advancements in chemistry End of chapter summaries review questions and exercises reinforce learning and facilitate self assessment Supplementary materials including online resources and laboratory exercises enhance the learning experience Whether you re a student seeking an introduction to applied chemistry or a professional looking to expand your knowledge Applied Chemistry Practical Applications is an invaluable resource for understanding the practical aspects of chemistry in industry technology and society

**Bionanocomposites** Carole Aimé, Thibaud Coradin, 2017-07-24 Beginning with a general overview of nanocomposites Bionanocomposites Integrating Biological Processes for Bio inspired Nanotechnologies details the systems available in nature nucleic acids proteins carbohydrates lipids that can be integrated within suitable inorganic matrices for specific applications Describing the relationship between architecture hierarchy and function this book aims at pointing out how bio systems can be key components of nanocomposites The text then reviews the design principles structures functions and applications of bionanocomposites It also includes a section presenting related technical methods to help readers identify and understand the most widely used analytical tools such as mass spectrometry calorimetry and impedance spectroscopy among others

Programmed Learning and Individually Paced Instruction Carl H. Hendershot,1973 PDE Toolbox Primer for Engineering Applications with MATLAB® Basics Leonid Burstein,2022-06-07 Partial differential equations PDEs describe technological phenomena and processes used for the analysis design and modeling of technical products Solutions of spatial and transient PDEs are realized by using the PDE Toolbox included in the MATLAB software MATLAB is introduced here as an essential foundation for PDE and the Modeler of the PDE Toolbox with appropriate explanatory solutions is

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