

Chapter 4: Supramolecular Problems - Answers

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a) CO_2 b) CH_4 c) HNO_3 d) Ca(OH)_2 e) H_2SO_4 f) Ag
 $\text{C} = +4$ $\text{C} = -4$ $\text{N} = +5$ $\text{Ca} = +2$ $\text{S} = +6$ $\text{Ag} = 0$
 $\text{O} = -2$ $\text{H} = +1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Ag} = 0$
 $\text{O} = -2$
- g) Fe h) BaCl_2 i) CO_2 j) H_2O k) HNO_3 l) Fe
 $\text{Fe} = 0$ $\text{Ba} = +2$ $\text{C} = +4$ $\text{H} = +1$ $\text{N} = +5$ $\text{Fe} = 0$
 $\text{Cl} = -1$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{Fe} = 0$
 $\text{Cl} = -1$

2. Assign the oxidation number to each atom in the following ions:

- a) NO^+ b) CO_3^{2-} c) HNO_2 d) SO_4^{2-} e) PO_4^{3-} f) HCO_3^-
 $\text{N} = +2$ $\text{C} = +3$ $\text{N} = +3$ $\text{S} = +6$ $\text{P} = +5$ $\text{C} = +3$
 $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$
 $\text{O} = -2$

3. For the following reactions identify if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidizing agent, the reducing agent, the substance being oxidized and the substance being reduced. Identify the oxidation number change to help to support these observations:



Redox rxn. Cu oxidized (oxidizing agent); H_2 as H^+ \longrightarrow H_2
 HCl reduced (oxidizing agent); H_2 as H^+ \longrightarrow H_2



Not a redox rxn. Ca as Ca^{2+} is both oxidized and reduced



Redox rxn. Fe oxidized (oxidizing agent); Si as $\text{Si}^0 \longrightarrow \text{Si}^+$
 Si reduced (oxidizing agent); Si as $\text{Si}^0 \longrightarrow \text{Si}^+$



Not a redox reaction. No change in Ag



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