



# Hp 81134a Performance Verification User Guide

**JA Banks**



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web the two general kinds of nuclear reactions are nuclear decay reactions and nuclear transmutation reactions in a nuclear decay reaction also called radioactive decay an unstable nucleus emits radiation and is transformed into the nucleus of one or more other elements the resulting daughter nuclei have a lower mass and are lower in energy

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web chapter 18 radioactivity and nuclear reactions 18 1 radioactivity 18 2 nuclear decay 18 3 detecting radioactivity 18 4 nuclear reactions radioactivity 18 1 the nucleus recall that atoms are composed of protons neutrons and electrons the nucleus of an atom contains the protons which have a positive charge and neutrons

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web sodium 23 in a sample of  $\text{NaCl}$  is subjected to neutron bombardment in a nuclear reactor to produce  $^{24}\text{Na}$  when removed from the reactor the sample is radioactive with beta activity of  $2.54 \times 10^4$  dpm the decrease in radioactivity over time was studied producing the following data

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web using algebra the moles was founded to be 0.00352 moles of butane in order to find the molarity the mass of the butane should be divided by the moles after further calculations the molarity of the butane gas was calculated to be 56.81 grams moles this lab is done with tons of sources of error

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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0.38 grams 0.02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp  $V_1/n_1 = V_2/n_2$   $0.0896 \text{ l} / n_2 = 22.4 \text{ l} / 1 \text{ mole}$   $n_2 = 0.0040 \text{ moles}$  12 experimental molar mass g mole of butane 0.23 g / 0.0040 moles 58 g mole 13 the molecular formula of butane is  $\text{C}_4\text{H}_{10}$

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web nov 3 2020 in this experiment the moles and mass will be required to determine the molar mass of butane in a butane lighter experimental to begin the experiment the mass of a butane lighter was measured before using the butane gas to deplete the water to the 80ml mark on a graduated cylinder

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web dec 18 2020 in this lab students will experimentally determine the molar mass of a gas specifically butane  $\text{C}_4\text{H}_{10}$  by collection over water this experiment is an inquiry based experiment for 2nd year chemistry or ap chemistry students who

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web the pressure of the gas the pressure of the gas was calculated using dalton s law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760mm hg the pressure of the water vapor was found by using the temperature of the water

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