

## Stoichiometry Practice 2.1 - Answer Key

**Objectives:** - Use stoichiometry to calculate the mass of any substance involved in a chemical reaction

**Directions:** Answer the following stoichiometry problems and show your work.

1. If a piece of magnesium with a mass of 2.76 g is added to a solution of hydrochloric acid (HCl), what mass of  $H_2$  would be produced? **[0.229 g]**



$$2.76 \text{ g Mg} \times \left( \frac{1 \text{ mol Mg}}{24.305 \text{ g Mg}} \right) \left( \frac{1 \text{ mol } H_2}{1 \text{ mol Mg}} \right) \left( \frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} \right) = \boxed{0.229 \text{ g } H_2}$$

2. What mass of barium chloride is needed to react completely with 46.8 g of sodium phosphate?



$$46.8 \text{ g } Na_3PO_4 \times \left( \frac{1 \text{ mol } Na_3PO_4}{163.94 \text{ g } Na_3PO_4} \right) \left( \frac{3 \text{ mol } BaCl_2}{2 \text{ mol } Na_3PO_4} \right) \left( \frac{208.23 \text{ g } BaCl_2}{1 \text{ mol } BaCl_2} \right) = \boxed{89.1 \text{ g } BaCl_2}$$

3. In the reaction, where magnesium oxide reacts with water to produce magnesium hydroxide. What mass of magnesium oxide is needed to produce 264 g of magnesium hydroxide?



$$264 \text{ g } Mg(OH)_2 \times \left( \frac{1 \text{ mol } Mg(OH)_2}{58.32 \text{ g } Mg(OH)_2} \right) \left( \frac{1 \text{ mol } MgO}{1 \text{ mol } Mg(OH)_2} \right) \left( \frac{40.30 \text{ g } MgO}{1 \text{ mol } MgO} \right) = \boxed{18.2 \text{ g } MgO}$$

4. Methane ( $CH_4$ ) burns in oxygen to produce carbon dioxide and water. What mass of oxygen would be necessary for the reaction if you had 15.0 g of methane?



$$15.0 \text{ g } CH_4 \times \left( \frac{1 \text{ mol } CH_4}{16.04 \text{ g } CH_4} \right) \left( \frac{2 \text{ mol } O_2}{1 \text{ mol } CH_4} \right) \left( \frac{32.00 \text{ g } O_2}{1 \text{ mol } O_2} \right) = \boxed{59.8 \text{ g } O_2}$$

# Mass Mass Stoichiometry Practice Key

**RC Schank**



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