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#### CHEMISTRY PAPER 2 (233/2)

#### FORM FOUR (4)

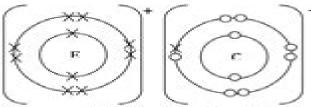
#### TIME: 2 HOURS

#### MARKING SCHEME

 $(1 \, \text{mk})$ (1, mik)

- 1 (a) Alkaline-earth metals
  - (b) B
  - (c) (i) A<sub>2</sub>B<sub>3</sub> (½mk)
- (ii) H<sub>2</sub>SO<sub>4</sub>
  - (3/2mk).





Distribution of electrons (1mk) and charges (1mk)

No label (0mk)

(e) D has a smaller atomic size than C//C has larger atomic size than D (1mk)

D has more protons than C//D has stronger nuclear charge than C (1mk)

(f) H (½mk). Loses electrons most readily/largest atomic radius/weakest nuclear charge (explanation ½mk)

(g) 2A(s) + 6HCl(aq) → 2ACl3(aq) + 3H2(g) 1mk

Moles of A = 
$$\frac{0.081}{27}$$
 = 0.003moles ½mk

Mole ratio 1:3

Moles of HCl = 0.003x3 = 0.009 moles  $\frac{1}{2}$ mk

0.009×1000 ½mk = 0.45M ½mk

(h) The melting point increase with increase in atomic number. 1mk

This is due to increase in the strength of van der waal's forces (inter-molecular forces) 1mk

# **Marking Scheme For Chemistry Paper2 2002**

Gabriell Ilevbare, Jeremy Busby, Peter Andersen

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