Worksheet: Writing and Balancing Chemical Reactions

- Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.
- a. $C_0 \hookrightarrow C_0 \hookrightarrow C$ **b.** ____ $H_2O_{(0)} \rightarrow$ ____ $H_{2((0)} +$ ____ $O_{2((0))}$ c. ____ Fe_(i) + ____ H₂O_(g) \rightarrow ____ H_{2(g)} + ____ Fe₃O_{4(g)} d. ____ AsCl_{3 (ee)} + ____ H₂S (ee) \rightarrow ____ As₂S_{3 (e)} + ____ HCl (ee) e. ____ CuSO₄ • 5 $H_2O_{(0)} \rightarrow$ ____ CuSO₄₍₀₎ + ____ $H_2O_{(0)}$ f. ____ Fe₂O_{3 ω 0 + ____ H₂ ω 0 \rightarrow ____ Fe ω 0 + ____ H₂O ω 0} g. $CaCO_{3 (n)} \rightarrow CaO_{(n)} + CO_{2 (n)}$ h. ____ Fe $_{(0)}$ + ____ S_{8 (0)} \rightarrow ____ FeS $_{(0)}$ i. ____ $H_2S_{(aq)} +$ ____ $KOH_{(aq)} \rightarrow$ ____ $H_2O_{(0)} +$ ____ $K_2S_{(aq)}$ j. ____ NaCl $_{(0)} \rightarrow$ ____ Na $_{(0)} +$ ____ Cl₂₍₀₎ k. ____ $Al_{(0)}$ + ____ $H_2SO_{4(aq)}$ \rightarrow ____ $H_{2(q)}$ + ____ $Al_2(SO_4)_{3(aq)}$ 1. _____ H_3PO_4 (oat) + _____ NH_4OH (oat) \rightarrow _____ H_2O (t) + _____ (NH_4) $_3PO_4$ (oat) m. ____ $C_3H_{8(g)} +$ ____ $O_{2(g)} \rightarrow$ ____ $CO_{2(g)} +$ ____ $H_2O_{(f)}$ o. ____ $CH_{4(g)} + ____ O_{2(g)} \rightarrow ____ CO_{2(g)} + ____ H_2O_{4h}$ p. ____ $K_2SO_{4 (aq)} +$ ____ $BaCl_{2 (aq)} \rightarrow$ ___ $KCl_{(aq)} +$ ___ $BaSO_{4 (a)}$ q. $C_5H_{12(0)} + C_2G_{00} \rightarrow CO_{2(6)} + H_{2O(6)}$ r. ____ Ca(OH)_{2 (aq)} + ____ NH₄Cl_(aq) \rightarrow ____ NH₄OH_(aq) + ____ CaCl_{2 (aq)} s. $\underline{\hspace{1cm}} V_2O_5\omega_1 + \underline{\hspace{1cm}} Ca\omega_2 \rightarrow \underline{\hspace{1cm}} CaO\omega_1 + \underline{\hspace{1cm}} V\omega_2$ t. $Na_{co} + Znl_{2coo} \rightarrow Nal_{coo} + Zn_{co}$ u. ____ $C_2H_6O_3 + ___ O_2 (y) \rightarrow ___ CO_2 (y) + ___ H_2O (y)$ w. ____ $Fe_2O_{3(x)} + ___ H_{2(x)} \rightarrow ___ Fe_{(x)} + ___ H_2O_{(0)}$ x. $C_{15}H_{30,00} + C_{2,(e)} \rightarrow C_{2,(e)} + H_{2}O_{60}$ y. ____ $BN_{(0)} + ____ F_{2(0)} \rightarrow ____ BF_{3(0)} + ____ N_{2(0)}$ z. $C_{12}H_{26(0)} + C_{2(0)} \rightarrow C_{2(0)} + H_{2}O_{(0)}$
- Predict the product(s) along with the states, indicate the type of reaction, and balance the following chemical reactions.
- a. A solution of lead (II) nitrate is mixed with a solution of sodium iodide.
- b. Solid zinc sulfide reacts with oxygen in the air.
- Liquid butane (C₄H_{10 (5)}) is used as a fuel to ignite a lighter.
- Barium hydroxide solution is neutralized by adding hydrochloric acid (HCl (app)).
- Copper metal is placed in a solution of silver nitrate.
- f. Sulfur burns in oxygen to make sulfur dioxide gas.
- g. A solution of aluminum sulfate is mixed with a solution of calcium hydroxide.
- Zinc metal is placed in sulfuric acid (H₂SO_{4 (aq)}).
- Aluminum powder is placed in a container filled with chlorine gas.
- Sucrose undergoes cellular respiration.

<u>Identifying And Balancing Chemical Equations Answers</u> Chapter 8

Padhraic Smyth

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