#### Section 7.2: Equilibrium Law and the Equilibrium Constant Tutorial 1 Practice, page 431

1. (a)  $2 \text{ CO}_2(g) \implies 2 \text{ CO}(g) + O_2(g)$ Products: 2 CO(g); O<sub>2</sub>(g) Reactant: 2 CO<sub>2</sub>(g) Equilibrium law equation:  $K = \frac{[CO(g)]^2[O_2(g)]}{[CO_2(g)]^2}$ (b)  $2 \text{ Cl}_2(g) + 2 \text{ H}_2O(g) \implies 4 \text{ HCl}(g) + O_2(g)$ Products: 4 HCl(g); O<sub>2</sub>(g) Reactants: 2 Cl<sub>2</sub>(g); 2 H<sub>2</sub>O(g) Equilibrium law equation:  $K = \frac{[HCl(g)]^4[O_2(g)]}{[Cl_2(g)]^2[H_2O(g)]^2}$ (c)  $2 O_0(g) \implies 3 O_2(g)$ Product:  $3 O_2(g)$ Reactant: 2 O<sub>y</sub>(g)

Equilibrium law equation:  $K = \frac{[O_2(g)]^3}{[O_2(g)]^2}$ 

(d)  $4 \text{ NH}_3(g) + 3 \text{ O}_2(g) \implies 2 \text{ N}_2(g) + 6 \text{ H}_2\text{O}(g)$ Products:  $2 N_2(g)$ ;  $6 H_2O(g)$ Reactants:  $4 \text{ NH}_3(g)$ ;  $3 \text{ O}_2(g)$ 

Equilibrium law equation:  $K = \frac{[N_2(g)]^2[H_2O(g)]^6}{[NH_2(g)]^4[O_2(g)]^3}$ 

 Given: CO(g) + 2 H<sub>2</sub>(g) = CH<sub>3</sub>OH(g); [CO(g)]<sub>equilibrium</sub> = 0.079 mol/L;  $[H_2(g)]_{\text{opathteium}} = 0.158 \text{ mol/L}; [CH_2OH(g)]_{\text{opathteium}} = 0.021 \text{ mol/L}.$ Required: K

Solution: Write the equilibrium law equation using the balanced chemical equation. Then, substitute the equilibrium concentrations into the equilibrium law equation and solve for K.

$$K = \frac{[CH_3OH(g)]}{[CO(g)][H_2(g)]^2}$$

$$= \frac{(0.021)}{(0.079)(0.158)^2}$$

K = 11

Statement: The equilibrium constant, K, for the reaction is 11.

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**Beatrice Mendez Newman** 

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