

Questions:

- (1) To identify the correct hydrocarbon structure corresponding to the following formula: $C_{10}H_{18}$
- (2) To calculate the number of moles of the starting material present in the solution.
- (3) To determine the limiting reagent.
- (4) To determine the theoretical amount of product that could be produced.
- (5) To compare the theoretical amount of product with the actual amount of product obtained.

Procedure:

- (1) Put on safety goggles.
- (2) Prepare a solution of the starting material in a beaker.
- (3) Put approximately 10 mL of sodium hydroxide solution in a graduated cylinder.
- (4) Measure the volume of the starting material in the beaker. Pour the sodium hydroxide solution into the beaker.
- (5) Carefully add the graduated cylinder into the beaker with sufficient speed. Stir the solution with a glass rod. Measure the volume of the starting material in the graduated cylinder. Repeat the addition four times. Pour the sodium hydroxide solution from the graduated cylinder into the beaker. Mix the solution with the glass rod. Measure the volume of the starting material in the graduated cylinder. Repeat the addition four times.
- (6) After adding the four solutions to the beaker, measure the volume of the solution in the beaker. Pour the solution into a beaker.
- (7) Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times.
- (8) After all of the solutions are added to the beaker, mix the solution with the glass rod. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times.
- (9) After all of the solutions are added to the beaker, mix the solution with the glass rod. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times.
- (10) After all of the solutions are added to the beaker, mix the solution with the glass rod. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times. Measure the volume of the starting material in the beaker. Repeat the addition four times.

State your observations:

Observations:	I have observed the reaction of the starting material with the sodium hydroxide solution. The reaction is exothermic and produces a precipitate.
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Essential Experiments For Chemistry 14d Questions

L Darling-Hammond



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