

c = specific heat ($J/g \cdot ^\circ C$) q = heat energy (J)
 T = temperature ($^\circ C$) m = mass (g) Δ = change in

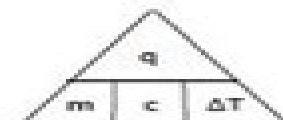
$$q = m \cdot c \cdot \Delta T$$

Name _____ Period _____

Thermal Energy & Heat Unit Study Guide

Part 1. For each of the following scenarios:

- Identify as vaporization, condensation, melting, or freezing
- Write the change in states [example: gas $\rightarrow$ liquid]
- Write whether the process absorbs or releases thermal energy



1. High altitude rain turns to hail.

a) _____ c) _____

b) _____

2. Towering cumulonimbus clouds form in the sky.

a) _____ c) _____

b) _____

3. A shallow puddle dries up over a couple day span.

a) _____ c) _____

b) _____

Part II. Specific Heat Calculations Refer to the equation and symbol legend above. Hint – use the triangle method. Show your work.

- A 110 g sample (m) of tin was heated from $23.2^\circ C$ to $39.8^\circ C$. What was the change in thermal energy? Solve for the heat energy (q). Tin has a specific heat (c) value of $0.2274 J/g \cdot ^\circ C$.

c = _____

m = _____

ΔT = _____

- A scientist heated bundled strands of copper wire from $21.0^\circ C$ to $82.0^\circ C$. Copper has a specific heat (c) capacity of $0.3845 J/g \cdot ^\circ C$. If the change in heat energy (q) was 192 J, what was the mass (m) of the bundled strands of copper wire?

c = _____

ΔT = _____

q = _____

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Clemens Wendtner



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