

Name _____ Date _____ Class _____

CHAPTER 16 STUDY GUIDE FOR CONTENT MASTERY**Energy and Chemical Change****Section 16.1 Energy**

In your notebook, read about the nature of energy.

In the space at the left, write true if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- True* _____ 1. Energy is the ability to do work or produce heat.
- cannot be* _____ 2. The law of conservation of energy states that energy can be created and destroyed.
- True* _____ 3. Chemical potential energy is energy stored in a substance because of its composition.
- True* _____ 4. Heat is a form of energy that flows from a warmer object to a colder object.
- True* _____ 5. A calorie is the amount of energy required to raise the temperature of one gram of pure water by one degree Celsius.
- False* _____ 6. A calorie is the SI unit of heat and energy.
- True* _____ 7. The specific heat of a substance is the amount of heat required to raise the temperature of one gram of that substance by one degree Celsius.

- True* _____ 8. Kinetic energy is energy of motion.
- both potential and kinetic energy* _____ 9. Chemical potential energy is a chemical reaction *contains only* potential energy.
- 1000 calories* _____ 10. One nutritional Calorie is equal to 100 calories.

- True* _____ 11. One calorie equals 4.184 joules.
- True* _____ 12. When a fuel is burned, some of its chemical potential energy is lost as heat.
- multiply* _____ 13. To convert kilojoules to joules, divide the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following questions. Show all your work.

14. If the temperature of a 500-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/g·°C.

$$q = c \times m \times \Delta T$$

$$q = 4.184 \text{ J/g}\cdot^{\circ}\text{C} \times 500 \text{ g} \times 2.00^{\circ}\text{C}$$

$$q = 4.184 \times 10^3 \text{ J} = 4.18 \text{ kJ}$$

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Chemistry: Matter and Change • Chapter 16

B1

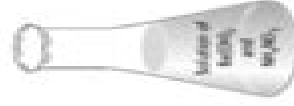
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CHAPTER 16 STUDY GUIDE FOR CONTENT MASTERY**Section 16.2 Heat in Chemical Reactions and Processes**

In your notebook, read about measuring heat and about chemical energy and the universe.

For each item in Column A, write the letter of the matching item in Column B.

- | Column A | Column B |
|--|--------------------------------|
| b. An insulated device used to measure the amount of heat absorbed or released during a chemical or physical process | a. system |
| c. The study of heat changes that accompany chemical reactions and phase changes | b. calorimeter |
| d. The specific part of the universe that contains the reaction or process you wish to study | c. thermodynamics |
| e. The change in enthalpy in a chemical reaction | d. universe |
| f. A system plus its surroundings | e. enthalpy |
| g. The heat content of a system at constant pressure | f. enthalpy change of reaction |
| h. Everything in the universe except the system being studied | g. surroundings |



Use the illustration to answer the following questions.

9. A solution is studying the solution in the flask. What is the system?
The system is the solution of NaOH and Na₃PO₄.
10. What are the surroundings?
The surroundings include everything except the solution.
11. What is the universe?
The universe is the solution plus the surroundings.

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Chemistry: Matter and Change • Chapter 16

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Energy And Chemical Change Study Guide Answer

Arihant Experts



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Student Study Guide/Solutions Manual for Essentials of General, Organic, and Biochemistry Denise Guinn, Rebecca Brewer, Rachel C. Lum, 2009-09-15 The Student Study Guide and Solutions Manual provides students with a combined manual designed to help them avoid common mistakes and understand key concepts. After a brief review of each section's critical ideas, students are taken through stepped-out worked examples, try-it-yourself examples, and chapter quizzes, all structured to reinforce chapter objectives and build problem-solving techniques. The solutions manual includes detailed solutions to all odd-numbered exercises in the text.

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon, buckyballs, graphite, and diamond, are illustrated at the left, as is the molecule methane (CH_4) from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications.

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