

Name Key Period _____

Worksheet: Calculating Empirical & Molecular Formulas

1. The empirical formula for the compound having the formula $H_4C_2O_4$ is:

- [A] C_2H_2 [B] CO_2H [C] COH [D] $C_2O_4H_4$ [E] COH_2

2. Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

- [A] CH_2 [B] CH [C] C_2H_4 [D] C_2H_6 [E] C_3H_8

$$C: \frac{85.6}{12.01} = 7.12 \quad H: \frac{14.4}{1.01} = 14.25 = \frac{14.25}{7.12} = 2$$

3. A compound is analyzed and found to contain 12.1% carbon, 56.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

- [A] $COCl_2$ [B] CO_2Cl_2 [C] CO_2Cl [D] $COCH$ [E] $COCl$

$$C: \frac{12.1}{12.01} = 1 \quad O: \frac{56.2}{16.00} = 3.51 \quad Cl: \frac{71.7}{35.45} = 2$$

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Calculate the empirical formula.

- [A] C_2H_2O [B] CH_2O [C] CH_4O [D] $C_2H_4O_2$ [E] C_2H_4O

$$C: \frac{40.0}{12.01} = 3.33 \quad H: \frac{6.7}{1.01} = 6.63 \quad O: \frac{53.3}{16.00} = 3.33 = \frac{3.33}{3.33} = 1$$

5. A compound contains 25.94% N and 74.06% O (by mass). What is the empirical formula?

$$\frac{25.94g N}{14.01} = \frac{1.85}{1.84} = 1 \times 2 = 2$$

$$\frac{74.06g O}{16.00} = \frac{4.63}{1.84} = 2.5 \times 2 = 5$$



6. Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

$$\frac{54.2g F}{19.00g/mol} = \frac{2.85}{1.43} mol F = 1.99 \sim 2$$

$$\frac{45.8g S}{32.07g/mol} = \frac{1.43}{1.43} mol S = 1$$



7. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

$$\frac{40.68g C}{12.01g/mol} = \frac{3.387}{3.387} \quad \frac{5.12g H}{1.01g/mol} = \frac{5.069}{3.387} \quad \frac{54.20g O}{16.00g/mol} = \frac{3.387}{3.387}$$

$$= 1 \times 2 = 2 \quad = 1.496 \times 2 = 3 \quad = 1 \times 2 = 2$$



8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

$$10.29g - 7.33g La = 2.96g O$$

$$\frac{7.33g La}{138.91g/mol} = .0527 \quad \frac{2.96g O}{16.00g/mol} = .1850 = \frac{.1850}{.0527} = 3.5 \times 2 = 7$$

$$= 1 \times 2 = 2$$



9. Calculate the molecular formula of a compound with the empirical formula CH_2O and a molar mass of 150 g/mol.

- [A] $C_7H_6O_3$ [B] $C_5H_{10}O_5$ [C] $C_3H_6O_3$ [D] $C_4H_8O_4$ [E] $C_6H_{12}O_6$

$$30 \quad 150 \div 30 = 5$$

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Table of Contents Empirical Formula Study Guide With Answer Sheet

1. Understanding the eBook Empirical Formula Study Guide With Answer Sheet
 - The Rise of Digital Reading Empirical Formula Study Guide With Answer Sheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Empirical Formula Study Guide With Answer Sheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Empirical Formula Study Guide With Answer Sheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Empirical Formula Study Guide With Answer Sheet
 - Personalized Recommendations
 - Empirical Formula Study Guide With Answer Sheet User Reviews and Ratings
 - Empirical Formula Study Guide With Answer Sheet and Bestseller Lists
5. Accessing Empirical Formula Study Guide With Answer Sheet Free and Paid eBooks
 - Empirical Formula Study Guide With Answer Sheet Public Domain eBooks
 - Empirical Formula Study Guide With Answer Sheet eBook Subscription Services
 - Empirical Formula Study Guide With Answer Sheet Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Empirical Formula Study Guide With Answer Sheet Compatibility with Devices
 - Empirical Formula Study Guide With Answer Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Empirical Formula Study Guide With Answer Sheet
 - Highlighting and Note-Taking Empirical Formula Study Guide With Answer Sheet
 - Interactive Elements Empirical Formula Study Guide With Answer Sheet
8. Staying Engaged with Empirical Formula Study Guide With Answer Sheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Empirical Formula Study Guide With Answer Sheet
9. Balancing eBooks and Physical Books Empirical Formula Study Guide With Answer Sheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Empirical Formula Study Guide With Answer Sheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Empirical Formula Study Guide With Answer Sheet
 - Setting Reading Goals Empirical Formula Study Guide With Answer Sheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Empirical Formula Study Guide With Answer Sheet
 - Fact-Checking eBook Content of Empirical Formula Study Guide With Answer Sheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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