

## Determining Empirical Formulas

|                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                     |         |                          |
|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------|--------------------------|
| 1) 75% carbon, 25% hydrogen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                         | Answer: | $\text{CH}_4$            |
| Work:<br>$\text{C } 75\% \rightarrow 75\text{g C} \left  \frac{1 \text{ mol}}{12.01\text{g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$<br>$\text{H } 25\% \rightarrow 25\text{g H} \left  \frac{1 \text{ mol H}}{1.01\text{g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{CH}_4</math></div>                                                                                                                                                            |         |                          |
| 2) 52.7% potassium, 47.3% chlorine                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                  | Answer: | $\text{KCl}$             |
| Work:<br>$\text{K } 52.7\% \rightarrow 52.7\text{g K} \left  \frac{1 \text{ mol K}}{39.10\text{g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$<br>$\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left  \frac{1 \text{ mol Cl}}{35.45\text{g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{KCl}</math></div>                                                                                                                                               |         |                          |
| 3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                 | Answer: | $\text{AlPO}_4$          |
| Work:<br>$\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left  \frac{1 \text{ mol Al}}{26.98\text{g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$<br>$\text{P } 25.4\% \rightarrow 25.4\text{g P} \left  \frac{1 \text{ mol P}}{30.97\text{g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$<br>$\text{O } 52.5\% \rightarrow 52.5\text{g O} \left  \frac{1 \text{ mol O}}{16.00\text{g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{AlPO}_4</math></div> |         |                          |
| 4) 13% magnesium, 87% bromine                                                                                                                                                                                                                                                                                                                                                                                                                                                                                                       | Answer: | $\text{MgBr}_2$          |
| Work:<br>$13\text{g Mg} \left  \frac{1 \text{ mol Mg}}{24.31\text{g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$<br>$87\text{g Br} \left  \frac{1 \text{ mol Br}}{79.90\text{g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{MgBr}_2</math></div>                                                                                                                                                                                                  |         |                          |
| 5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen                                                                                                                                                                                                                                                                                                                                                                                                                                                                                         | Answer: | $\text{Na}_2\text{SO}_4$ |
| Work:<br>$32.4\text{g Na} \left  \frac{1 \text{ mol}}{22.99\text{g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$<br>$22.5\text{g S} \left  \frac{1 \text{ mol S}}{32.37\text{g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$<br>$45.1\text{g O} \left  \frac{1 \text{ mol O}}{16\text{g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{Na}_2\text{SO}_4</math></div>                                                                                    |         |                          |

# Empirical Formula Chemistry If8766 Answers

**M Carnoy**



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