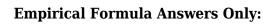
Example 1: Calculate the empirical formula for a substance containing 1,11 g of Se and 1.61 g of F.

1.11 g Se x
$$\frac{1 \text{ mol Se}}{78.971 \text{ g Se}} = 0.0141 \text{ mol Se}$$

1.16 g F x $\frac{1 \text{ mol f}}{18.998 \text{ g F}} = 0.0847 \text{ mol F}$
Se $\frac{0.0141}{0.0141} = \frac{5e}{0.0847} = \frac{5e}{6}$

Empirical Formula Answers Only

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