Date

Chemistry — Unit 5 Worksheet 3 Empirical and Molecular Formulas

Show all your work when solving the following problems. Be sure to include units and label your answer.

 Find the empirical formula of a compound containing 32.0 g of bromine and 4.9 g of magnesium.

32.0 g × <u>1 mole</u> = 0.401 mole Br 79.9 g

1:2 ratio

4.9 g × 1 mole = 0.20 mole Mg 24.31 g

MgBrz

What is the empirical formula of a carbon-oxygen compound, given that a 95.2 g sample of the compound contains 40.8 g of carbon and the rest oxygen?
95.2 g - 40.8 g = 54.4 g O

 $40.8 \text{ g} \times 1 \text{ mole} = 3.40 \text{ moles C}$ 12.0 g

1:1 ratio

54.4 g x× <u>1 mole</u> = 3.40 moles O 16.0g CO

3. A compound was analyzed and found to contain 9.8 g of nitrogen, 0.70 g of hydrogen, and 33.6 g of oxygen. What is the empirical formula of the compound?

9.8 g × 1 mole = 0.70 moles N 14.0g

1:1:3 ratio

0.70 g × <u>1 mole</u> = 0.69 moles H 1.01 g HNO₃

33.6 g × 1 mole = 2.10 moles O 16.0 g

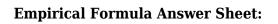
 A compound composed of hydrogen and oxygen is found to contain 0.59 g of hydrogen and 9.40 g of oxygen. The molar mass of this compound is 34.0 g/mol. Find the empirical and molecular formulas.

0.59 g × 1 mole = 0.58 mole H 1.01 g 9.40 g × 1 mole = 0.588 mole O 1:1 ratio; empirical formula: HO empirical formula mass: 17.0 g/mol molecular formula mass: 34.0 g/mol mfm is 2x efm so the molecular formula is: H-O-

16.0 g

Empirical Formula Answer Sheet

Richard Post, Chad Snyder, Clifford C. Houk



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