

Chapter 5 (The Periodic Table) Notes

- Sec 5.6 (Blocks and Sublevels)
 - We can use the periodic table to predict which sublevel is being filled by a particular element.
 - Refer to the periodic table and state the highest energy sublevel for each of the following elements:
 - He – 1s
 - K – 1s
 - U – 5f
 - Pd – 4d
 - Be – 2s
 - Co – 3d
 - Si – 3p
 - Pt – 5d
 - The electron configuration for Na is: $1s^2 2s^2 2p^6 3s^1$
 - We can abbreviate the electron configuration by indicating the innermost electrons with the symbol of the preceding noble gas.
 - The preceding noble gas with an atomic number less than sodium is neon, Ne.
- Sec 5.7 (Valence Electrons)
 - When an atom undergoes a chemical reaction, only the outermost electrons are involved.
 - These electrons are of the highest energy and are furthest away from the nucleus. These are the valence electrons.
 - The roman numeral in the American convention indicates the number of the valence electrons.
 - Group IA elements have one valence electron.
 - Group VA elements have five valence electrons.
 - When using the IUPAC designations for group numbers the last digit indicates the number of valence electrons.
 - Group 14 elements have four valence electrons.
 - Group 2 elements have two valence electrons.
- Sec 5.8 (Electron Dot Formulas)
 - An **electron dot formula** of an element shows that symbol of the element surrounded by its valence electrons.
 - We use one dot for each valence electron.

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