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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{CuCl}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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net ionic equation $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2\text{g}$ here is a video on spectator ions
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web the six NO_3^- ions and the six Na^+ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3\text{aq} + \text{KClaq} \rightarrow \text{AgCl s} + \text{KNO}_3\text{aq}$ total ionic $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$ net ionic $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s}$ 2

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{l}$ the spectator ions in KOH aq and HNO_3aq are K^+aq and NO_3^-aq how to identify and list the

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loss or gain of one or more electrons ions containing positive charge are cations and those containing negative charge are anions an ion with a net charge of zero is electrically neutral

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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