Empirical and Molecular Formulas Worksheet



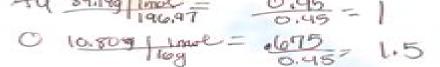
Objectives:

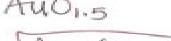
· be able to calculate empirical and molecular formulas

Empirical Formula

1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.521g of Oxygen?

2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O?





3) What is empirical formula if compound consists of 21.2%N, 6.1%H, 24.2%S and 48.5%O?

N 2129 | 1md - 1.5 . 2

H 19 | 1 md = 6.03 - 8

S 24.29 | 3.03 - 9

US.59 | 1 md | 3.03 - 4

- Molecular Formula J
- 4) Empirical formula of a substance is CH2O. Molar mass is 180. What is the molecular formula?



5) Sample (3.585g) contains 1.388g of C, 0.345g of H, 1.850g O and its molar mass is 62g. What is molecular formula of this substance?

Empirical Formula Practice With Answers

Julia Schneider

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