

# Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \times \frac{1 \text{ mol}}{12.01 \text{ g mol}^{-1}} \text{C} = 3.33 \text{ mol C}$$

$$\text{H}_{6.67} \times \frac{1 \text{ mol}}{1.008 \text{ g mol}^{-1}} \text{H} = 6.61 \text{ mol H}$$

C<sub>3</sub>H<sub>5</sub>

2. A compound is found to contain 20.0 g of carbon, 2.67 g of hydrogen, and 6.67 g of oxygen. Find the empirical formula.

$$\text{C}_{20.0} \text{H}_{2.67} \text{O}_{6.67} \times \frac{1 \text{ mol}}{12.01 \text{ g mol}^{-1}} \text{C} = 1.67 \text{ mol C}$$

$$\text{H}_{2.67} \times \frac{1 \text{ mol}}{1.008 \text{ g mol}^{-1}} \text{H} = 2.65 \text{ mol H}$$



$$\text{O}_{6.67} \times \frac{1 \text{ mol}}{16.00 \text{ g mol}^{-1}} \text{O} = 0.417 \text{ mol O} \quad \text{C}_2\text{H}_2\text{O}_4$$

3. A compound is found to contain 40.0 g of carbon, 6.67 g of hydrogen, and 13.33 g of oxygen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \text{O}_{13.33} \times \frac{1 \text{ mol}}{12.01 \text{ g mol}^{-1}} \text{C} = 3.33 \text{ mol C}$$



$$\text{H}_{6.67} \times \frac{1 \text{ mol}}{1.008 \text{ g mol}^{-1}} \text{H} = 6.61 \text{ mol H}$$

$$\text{O}_{13.33} \times \frac{1 \text{ mol}}{16.00 \text{ g mol}^{-1}} \text{O} = 0.833 \text{ mol O}$$

4. A 100.0 g compound is analyzed and found to contain 40.0 g carbon, 6.67 g hydrogen, and 53.33 g oxygen. Find the empirical formula.

$$\text{C}_{40.0} \text{H}_{6.67} \text{O}_{53.33} \times \frac{1 \text{ mol}}{12.01 \text{ g mol}^{-1}} \text{C} = 3.33 \text{ mol C}$$



$$\text{H}_{6.67} \times \frac{1 \text{ mol}}{1.008 \text{ g mol}^{-1}} \text{H} = 6.61 \text{ mol H}$$

$$\text{O}_{53.33} \times \frac{1 \text{ mol}}{16.00 \text{ g mol}^{-1}} \text{O} = 3.33 \text{ mol O}$$

# Empirical Formula Pogil

**Anthony S. Fauci**



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